

② 17.5 mde , ~~2.4M~~ , Find  $\rho$

2.4 moles  
 $\rho$

|                    |          |
|--------------------|----------|
| $\rho$             | 17.5 mde |
| <del>2.4 mde</del> |          |

= 7.29  $\rho$

⊕

$$\frac{M}{1} = \frac{\text{moles}}{\rho}$$


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$$\frac{2.4}{1} = \frac{17.5}{\rho}$$


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$$\rho = \frac{17.5}{2.4}$$

Apr 1-8:38 AM

② Find  $m = \frac{\text{moles solute}}{\text{Kg solvent}}$  12.5g NaOH

100g H<sub>2</sub>O

$$\frac{0.3125 \text{ mde}}{0.1 \text{ Kg}}$$

3.125M

|            |             |             |
|------------|-------------|-------------|
| 12.5g NaOH | 1 mole NaOH | = 3.125 mde |
|            | 40g NaOH    |             |

Apr 1-8:58 AM

(23)  $M = \frac{\text{moles}}{L}$

|  |  |
|--|--|
| <del>4.9 g H<sub>2</sub>SO<sub>4</sub></del> | 1 mole H <sub>2</sub> SO <sub>4</sub>        |
| 1 L soln                                     | <del>9.8 g H<sub>2</sub>SO<sub>4</sub></del> |

= 0.05 M

Apr 1-9:00 AM

(24) — g NaBr, ~~103 g NaBr~~, 30, ~~4 M~~

~~$\frac{1 \text{ mole NaBr}}{103 \text{ g NaBr}}$~~

~~$\frac{1 \text{ mole NaBr}}{12}$~~

|                        |                        |               |
|------------------------|------------------------|---------------|
| <del>103 g NaBr</del>  | <del>4 mole NaBr</del> | <del>30</del> |
| <del>1 mole NaBr</del> | <del>1</del>           | <del>12</del> |

= 1236 g NaBr

Apr 1-9:02 AM

(EC1)

$$m_{\text{de start}} = m_{\text{de end}}$$

$$M \times Q = M \times Q$$

$$\frac{(0.8)(100\text{mg})}{500} = \frac{M \times 500}{500}$$

0.16M

Apr 1-9:08 AM

(2) 58g NaCl, 5.6 mole H<sub>2</sub>O,

Solute                      Solvent

|          |             |  |  |
|----------|-------------|--|--|
| 58g NaCl | 1 mole NaCl |  |  |
| 58g NaCl | 58g NaCl    |  |  |

1 mole NaCl

|                           |                      |  |  |
|---------------------------|----------------------|--|--|
| 5.6 mole H <sub>2</sub> O | 18g/H <sub>2</sub> O |  |  |
| 1 mole H <sub>2</sub> O   | 18g/H <sub>2</sub> O |  |  |

M =  $\frac{\text{Mols solute}}{\text{kg Solvent}}$

0.1008

100.8 g H<sub>2</sub>O

0.1008 Kg

Apr 1-9:10 AM

$$\textcircled{3} \Delta T = (K \times \underline{m}) \cdot i$$

$$1.56 = (0.52 \times m) \cdot 1$$

$$\textcircled{3m}$$

$$\frac{101.56^\circ \text{SOLN}}{100 \text{ normal}}$$
$$\Delta T = 1.56$$

Apr 1-9:12 AM