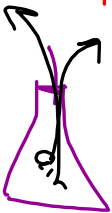


$T = 100^\circ\text{C}$
 $T = \text{Sink H}_2\text{O}$

Charles' Law $\frac{V}{T} = \text{constant.}$

$P_{\text{atm}} = 30.03 \text{ in}$ 2/24

TOTAL Volume of "air" in flask = $\sim 150 \text{ ml}$ ^{TOTAL}
 fill with water.
 Heat flask to 100°C



Air That left = $\sim 90 \text{ ml}$
 = H_2O in (sink).

Air left = ? 60 ml