

3/14 P.S.

(25) 50 ml gas $20^{\circ}\text{C} \rightarrow 30^{\circ}\text{C}$
 new V = ?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{303}{1} * \frac{50}{293} = \frac{V_2}{303} * \frac{303}{1}$$

(26) $\frac{0.5 \text{ mole}}{11.2 \text{ l}} = \frac{1 \text{ mole}}{22.4 \text{ l}}$
 STP

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(27) $\frac{PV}{T} = \frac{PV}{T}$

$$\frac{(101.3)(1400)}{20} = \frac{(50.65)V_2}{40}$$

(28) 7 mole A + 3 mole B $P = 1 \text{ atm}$

$$P_A = X_A (P_T) = \frac{7}{10} (1) = 0.7 \text{ atm}$$

(40)

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LAB - BP ↑

Add solute to a solvent

① FP ↓
② BP ↑
③ osmotic pressure

$\Delta T = (K \times m) i$

↑ valicity ↑ #ions

NaCl ②
CaCl₂ ③
C₆H₁₂O₆ ①

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NaCl → Na⁺ + Cl⁻

① (1 + 1) ②

CaCl₂ → Ca⁺² + 2Cl⁻

 1 + 2 = ③

C₆H₁₂O₆ → N/A (M) ① molecules

(solvent)

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metal NaCl soln

$$\rightarrow \frac{\text{moles NaCl}}{\text{kg H}_2\text{O}}$$

$$1 \text{ kg} \sim 1 \text{ l}$$

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