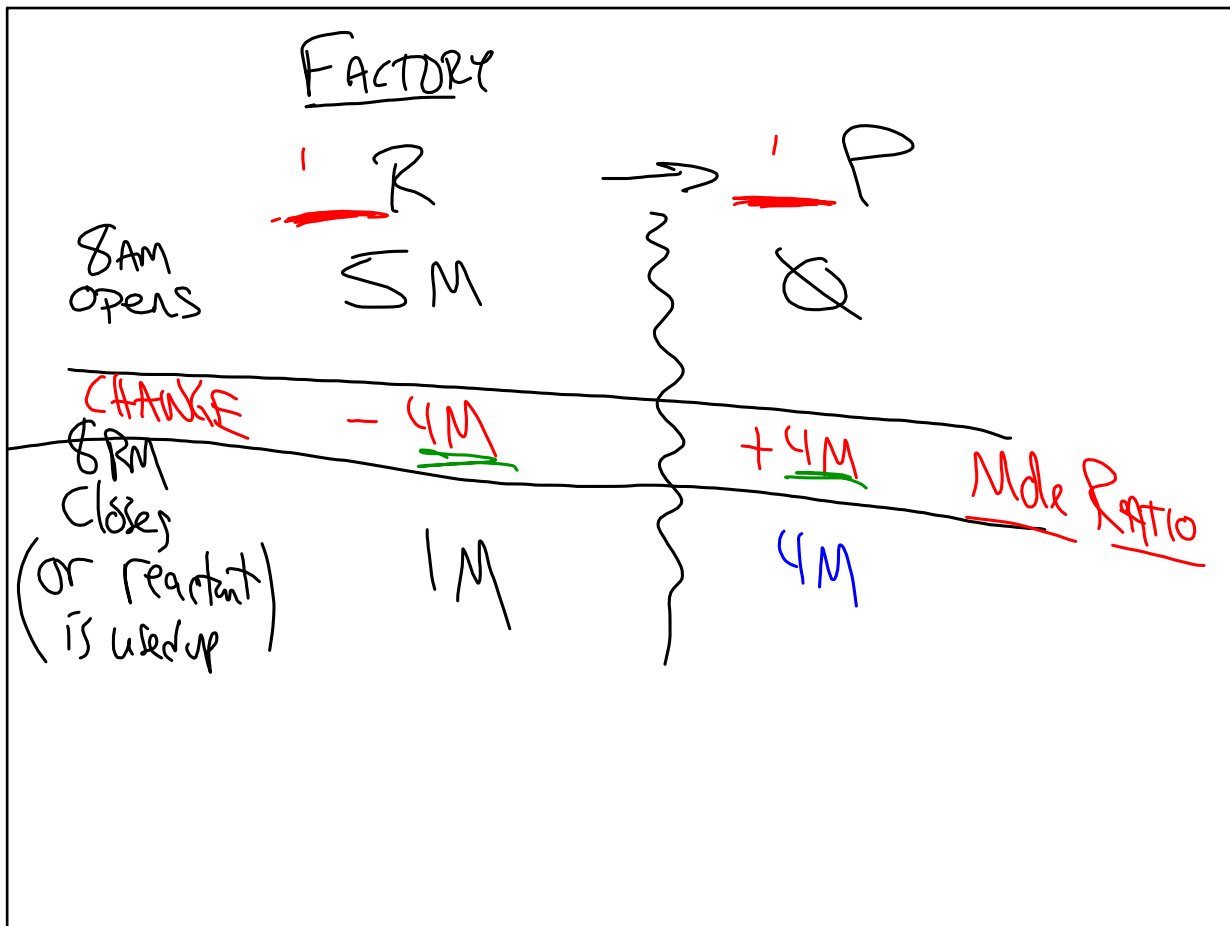


Apr 5-9:28 AM



Apr 5-9:29 AM

FACTORY #2

Reactants 1 : 2 mole ratio Products

time →

	<u>1 L</u>	→	<u>2R</u>
initial	5M		0
<u>Change</u>	<u>-4M</u>		<u>+8M</u>
end	1M		8M

MOLE RATIO

Rate of Change $\frac{M}{sec}$

$\frac{\Delta[M]}{\Delta t}$

↑
concentration of [H₂O]

Apr 5-9:32 AM

investigate the rate of a rxn

$2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$

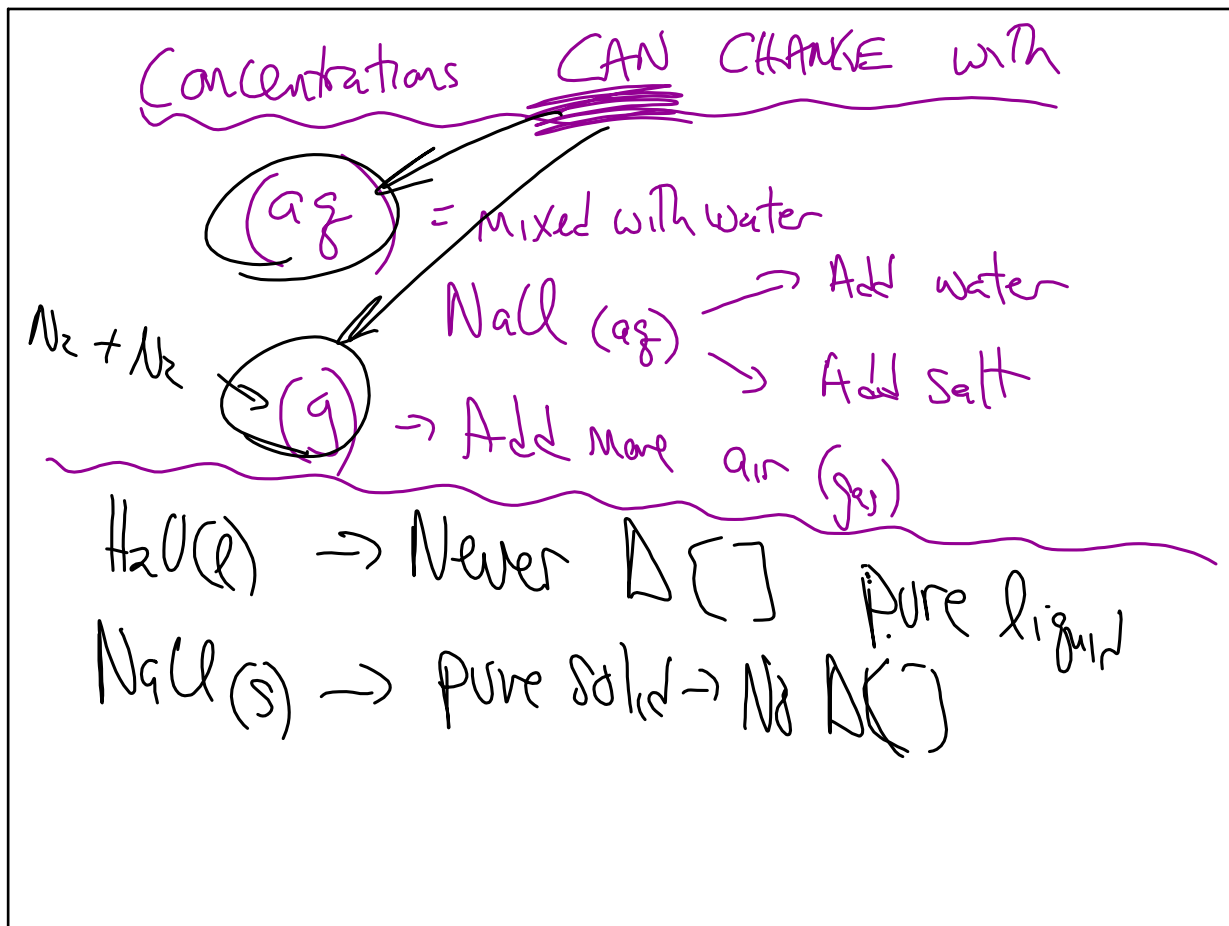
Equilibrium constant (K)

$$K = \frac{[\text{Products}]^{\text{coeff.}}}{[\text{Reactants}]^{\text{coeff.}}} = \frac{(\text{Product})^{\text{coeff.}}}{(\text{React})^{\text{coeff.}}}$$

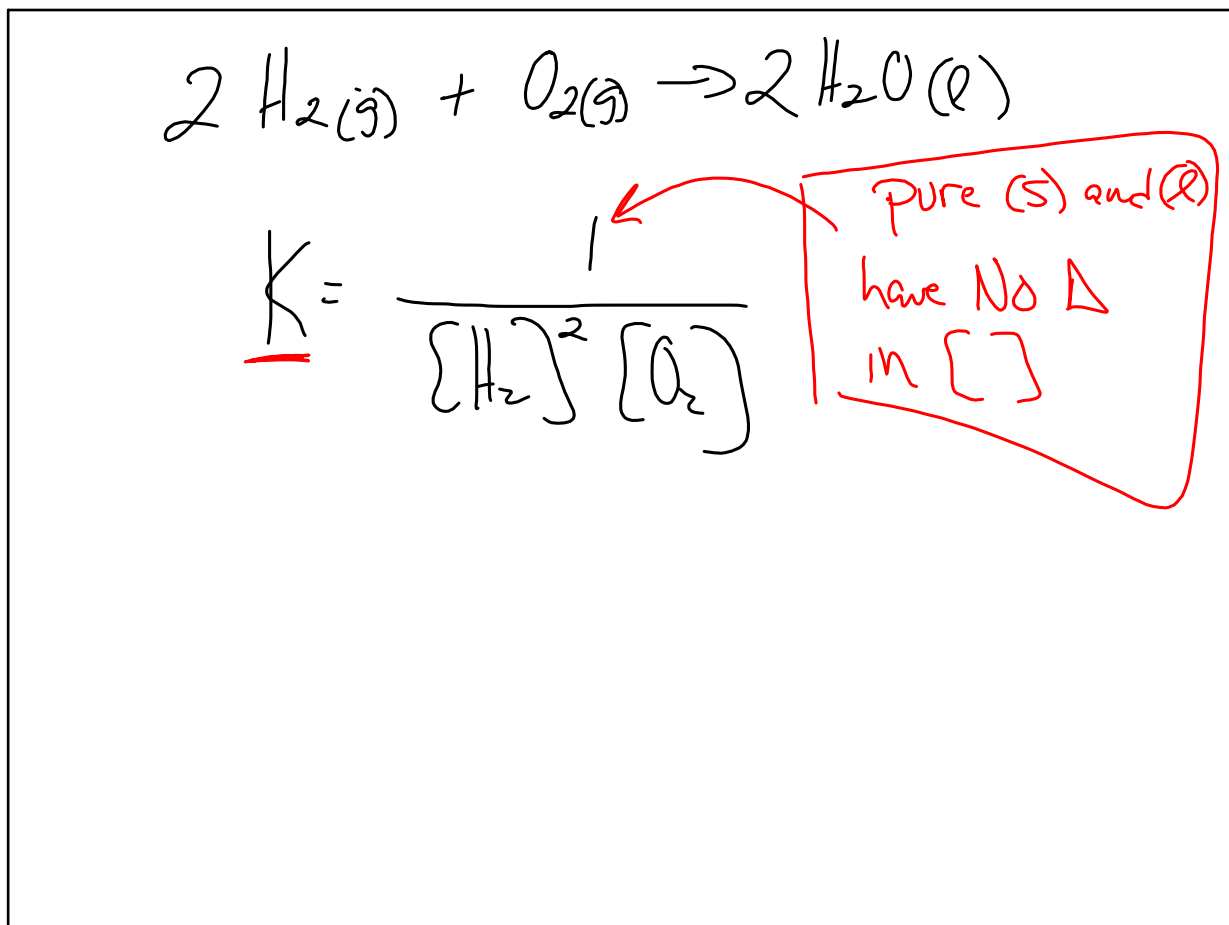
$$K = \frac{[H_2O]^2}{[H_2]^2 [O_2]}$$

Equilibrium constant expression.

Apr 5-9:41 AM



Apr 5-9:47 AM



Apr 5-9:52 AM

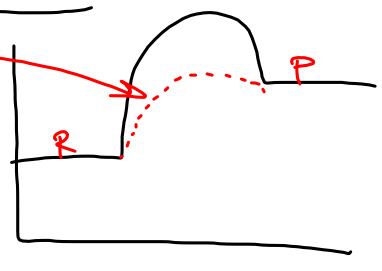
Speedup the rate of rxn

① Add Catalyst

② $\uparrow T$

② $\uparrow P$ (gas)

③ Move it \rightarrow STIR
Shake



Speedup Molecules
More collisions.

Apr 5-9:53 AM

HW P / 1, 2, 15

\uparrow K expression
(s) or (l) = "1"

Solubility curve Lab

Apr 5-9:57 AM