

May 5-9:26 AM

$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$

IONS NEVER GET SUBSCRIPTS

Half Rxn: $\text{H}_2 \rightarrow 2\text{H}^{+1} + 2\text{e}^{-}$
 $\text{O}_2 + 4\text{e}^{-} \rightarrow 2\text{O}^{-2}$

① Balanced elements.
 ② Balance the charges by ADDING e⁻s

gain to reactant, lose from product

$2(\text{H}_2 \rightarrow 2\text{H}^{+1} + 2\text{e}^{-})$
 $\text{O}_2 + 4\text{e}^{-} \rightarrow 2\text{O}^{-2}$

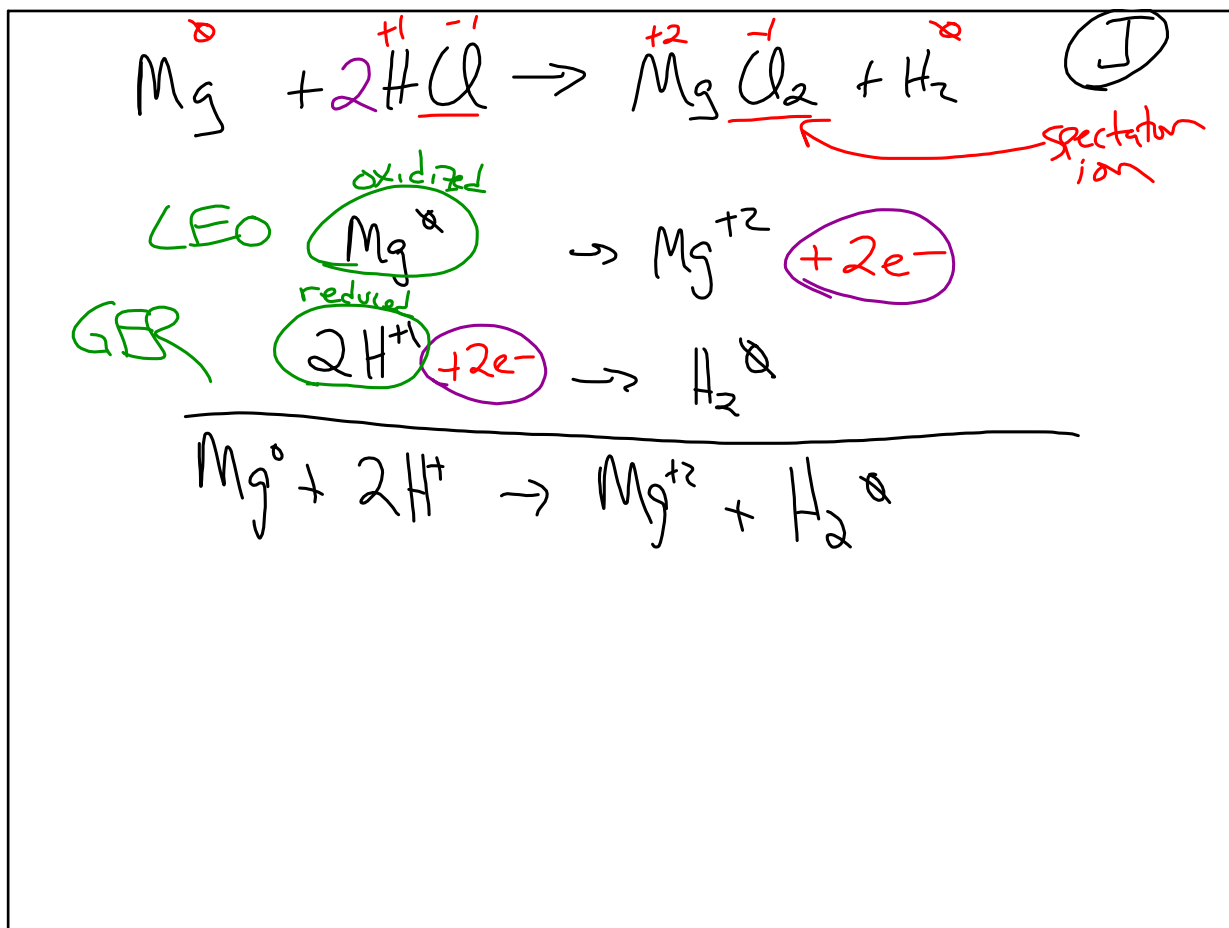
$2\text{H}_2 \rightarrow 4\text{H}^{+1} + \cancel{4\text{e}^{-}}$
 $\text{O}_2 + \cancel{4\text{e}^{-}} \rightarrow 2\text{O}^{-2}$

$2\text{H}_2 + \text{O}_2 \rightarrow 4\text{H}^{+1} + 2\text{O}^{-2}$

Used REDOX to balance the eqn

③ e⁻ gained = e⁻ lost
 mult by L.C.F.
 (S.C.M)

May 5-9:32 AM



May 5-9:53 AM

HW p 9 + 10

(AR)

May 5-10:03 AM