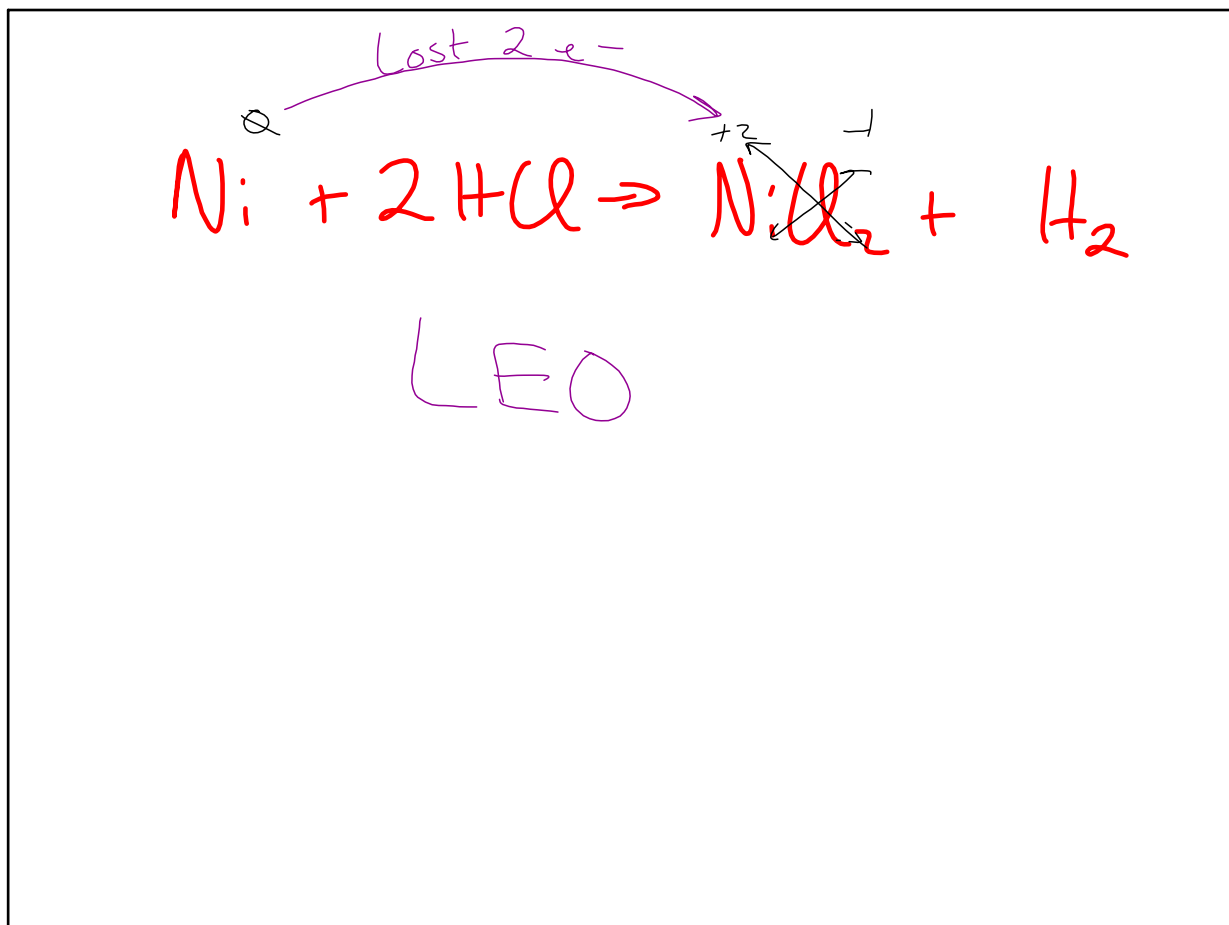


(At. mass)  $\rightarrow$  42  
 #P + #N  $\rightarrow$  19 K  
 #P  $\rightarrow$  19  
 (At #)  $\rightarrow$  23n

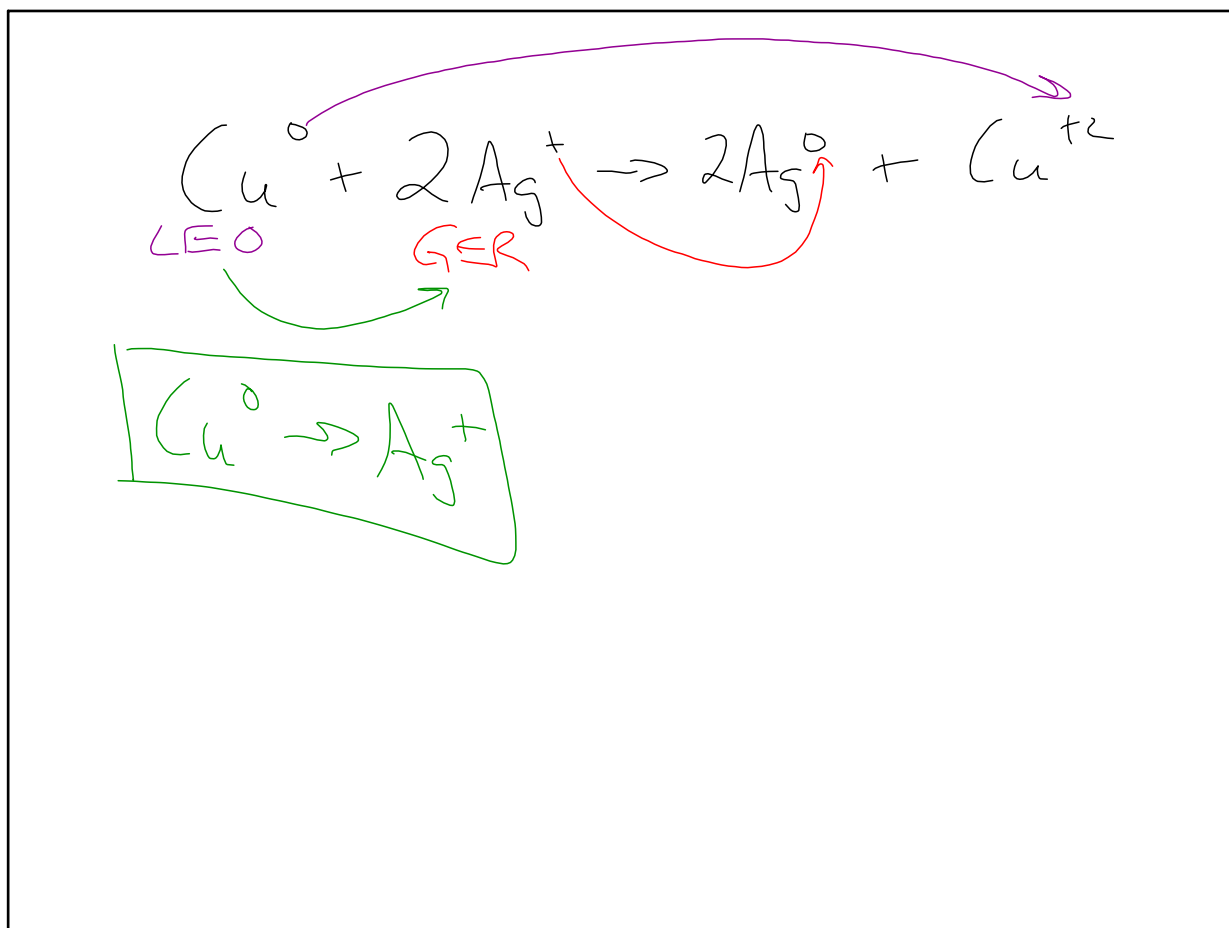
1.5 mder, 500 ml  
 (0.5 l)

$$\frac{M}{l} = \frac{\text{Moles}}{l} = \frac{1.5 \text{ mole}}{0.5 l} = 3.0$$

Aug 14-10:00 AM



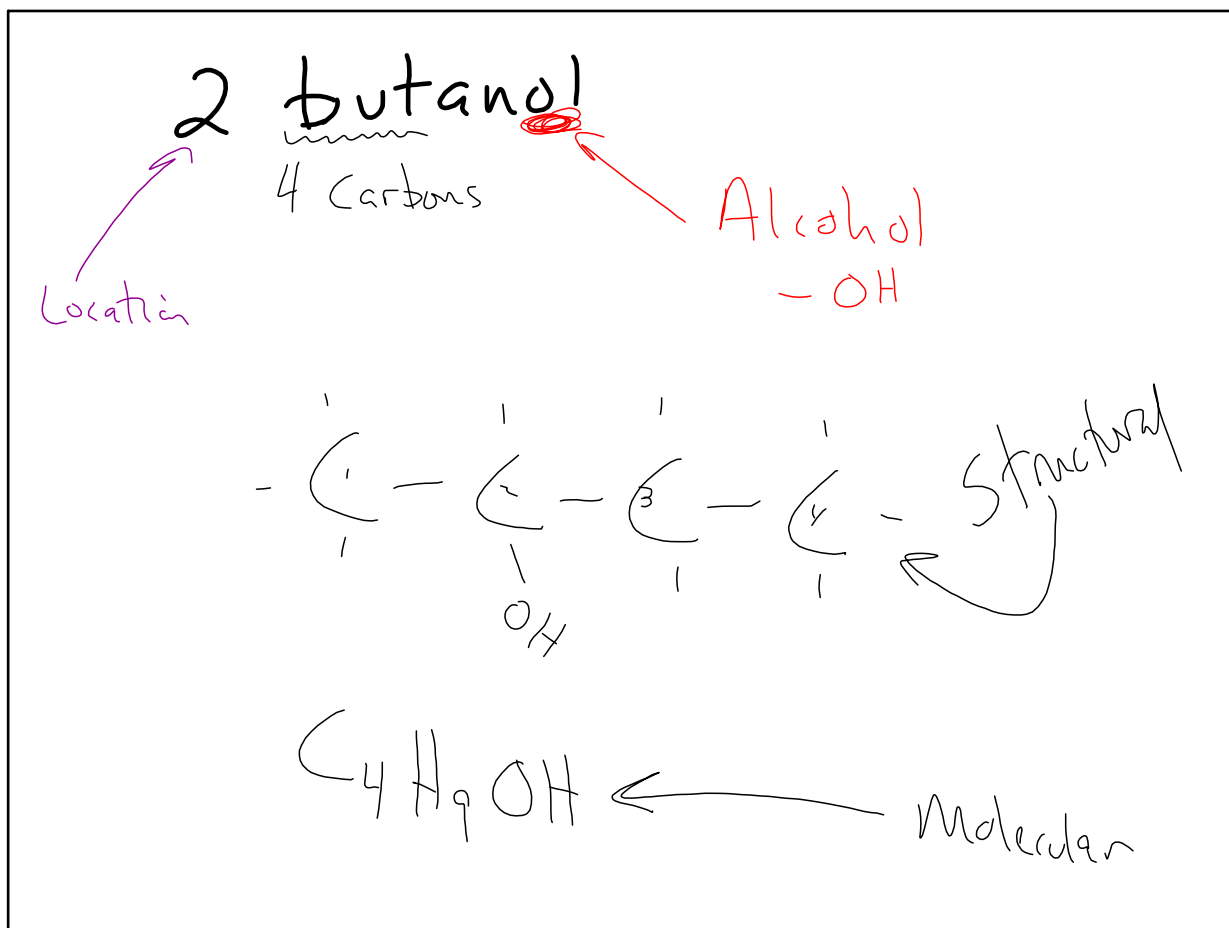
Aug 14-10:45 AM



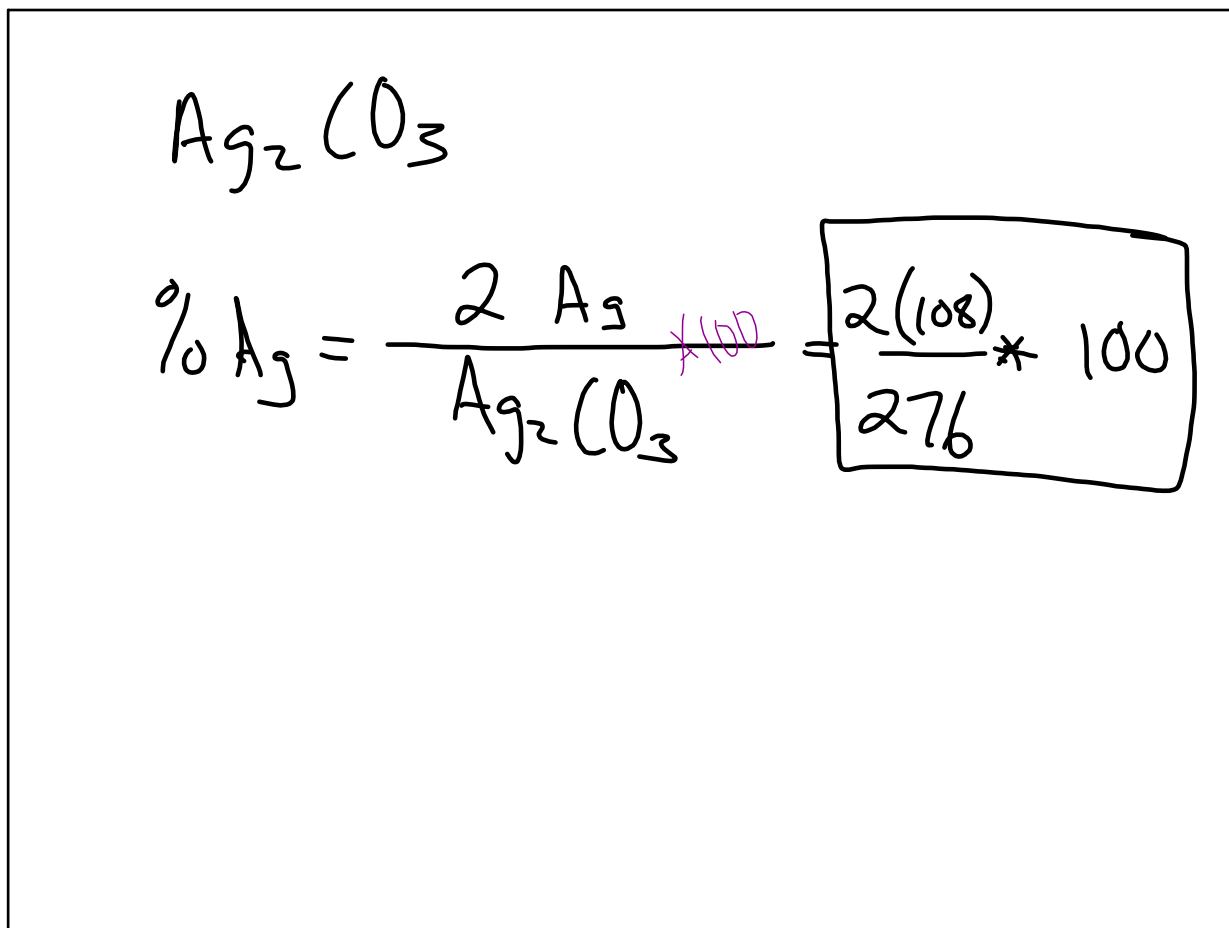
Aug 14-10:50 AM

	<u>Acids</u>	<u>Bases</u>
Arrhenius	Have $\text{H}^+$ ONLY $\oplus$ ion	has $\text{OH}^-$ ONLY $\ominus$ ion
Bronsted Lowry	Donates $\text{H}^+$ gives $\rightarrow$	Accept take $\text{H}^+$

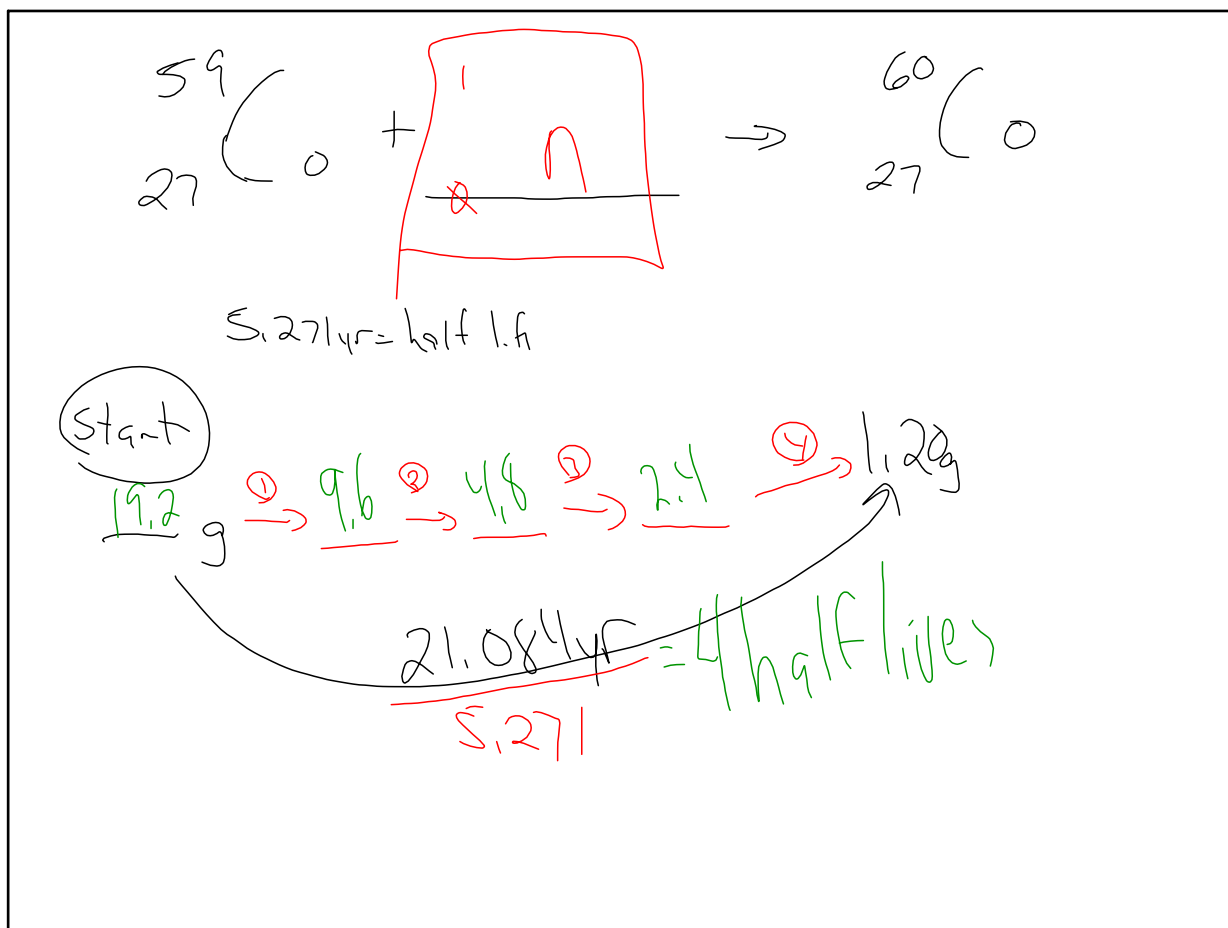
Aug 14-10:53 AM



Aug 14-11:17 AM



Aug 14-11:23 AM

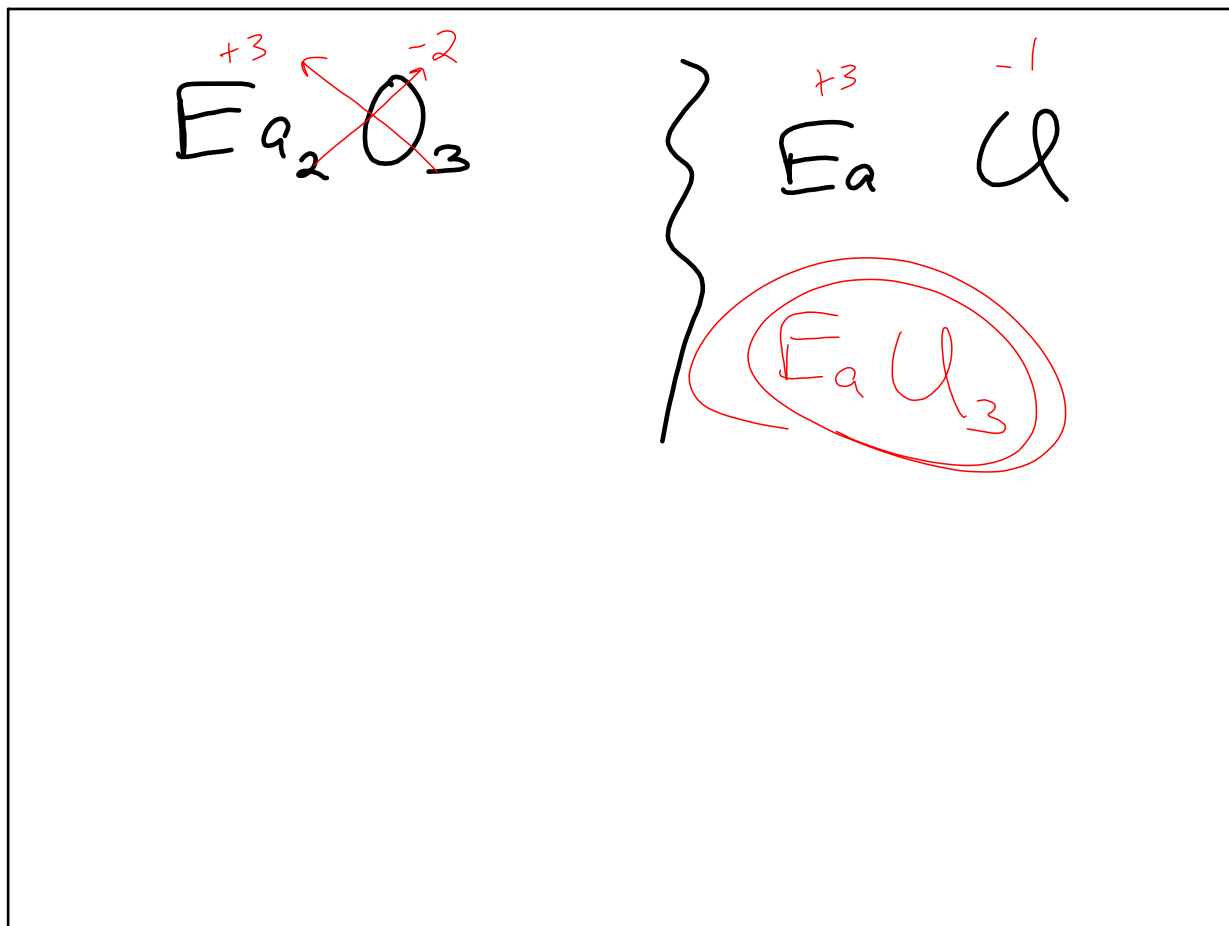


Aug 14-11:28 AM

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{(0.5)(24.5)}{298} = \frac{1(V_2)}{265}$$

Aug 14-11:43 AM



Aug 14-11:48 AM