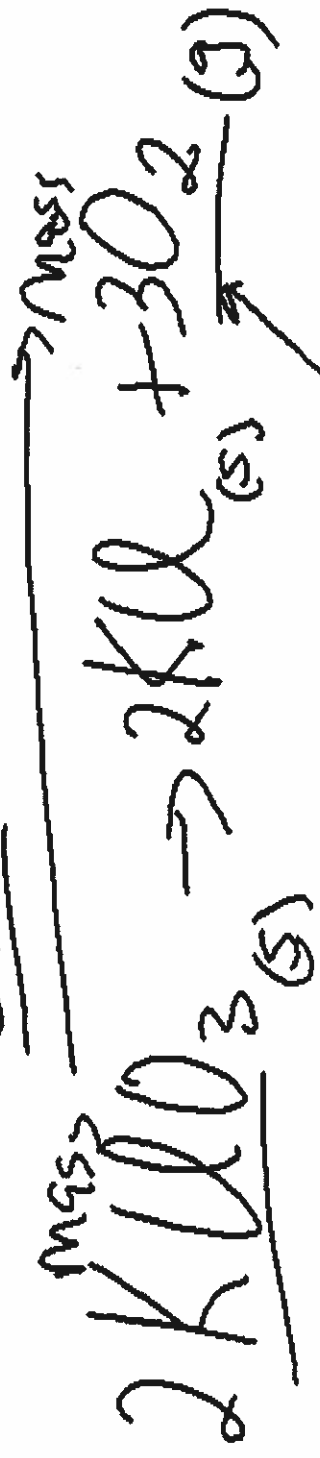


LAB



density

① % Oxygen = $\frac{\text{Mass lost R} \rightarrow \text{P (oxygen)}}{\text{KClO}_3 (3g)}$

② Calculated Actual % of O₂ we should get (F.W) from 9 KClO₃ we started with

density

3g KClO ₃	3 mole KClO ₃	3 mole O ₂	3g O ₂
122.5g KClO ₃	2nd KClO ₃	Indefinite	Indefinite

