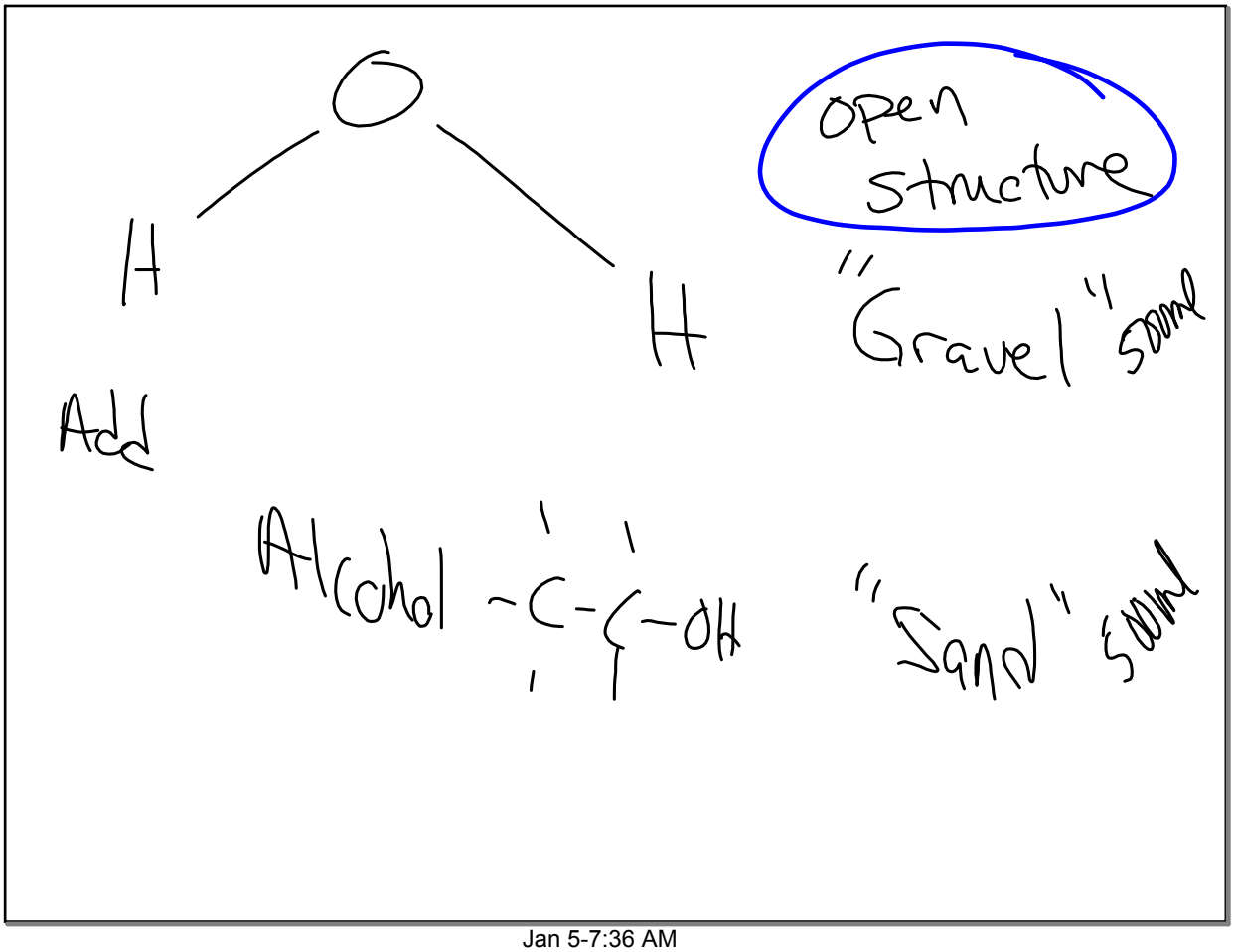
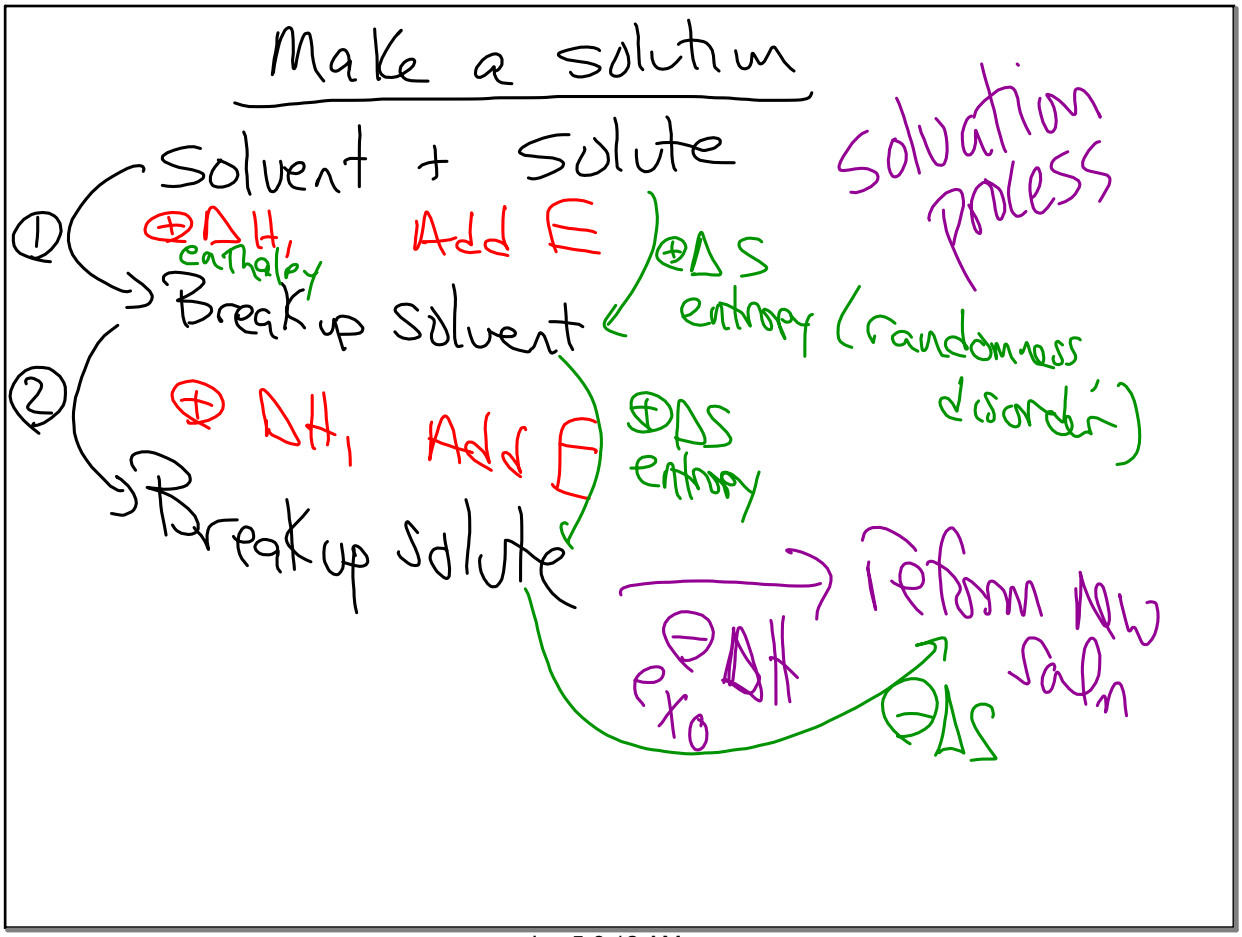


Jan 5-7:36 AM



Jan 5-8:12 AM

Solution Types

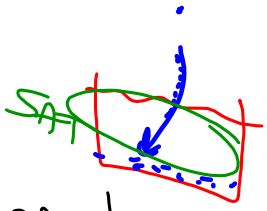
① Unsaturated

"Solvent can hold more solute

"Vacancy" at the Inn

② Saturated

Solvent holds as much solute as it can.



Jan 5-8:18 AM

③ Supersaturated "SUPA-SAT"

Solvent is holding MORE solute Than it can under std. conditions

SPECIAL CONDITIONS!

↑P, ↑T, ↑V

(gas)

Jan 5-8:24 AM

Miscible - mix

immiscible - don't mix

Factors that affect solubility

1. Bonding
2. T
3. P (g)
4. Concentration already Present

Jan 5-8:50 AM

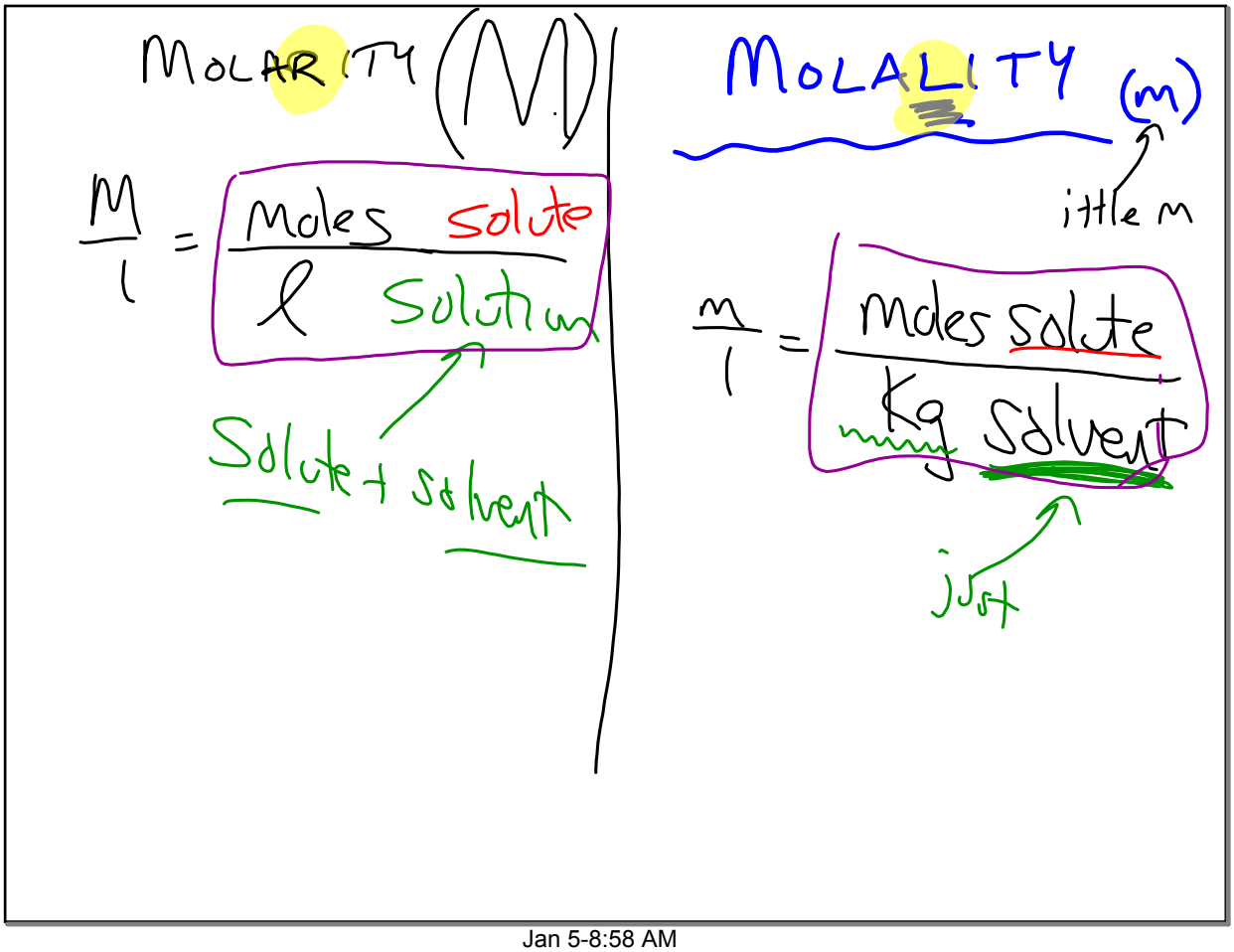
MOLARITY measure of concentration

$$\frac{M}{1} = \frac{\text{moles of solute}}{\text{l of solution}}$$

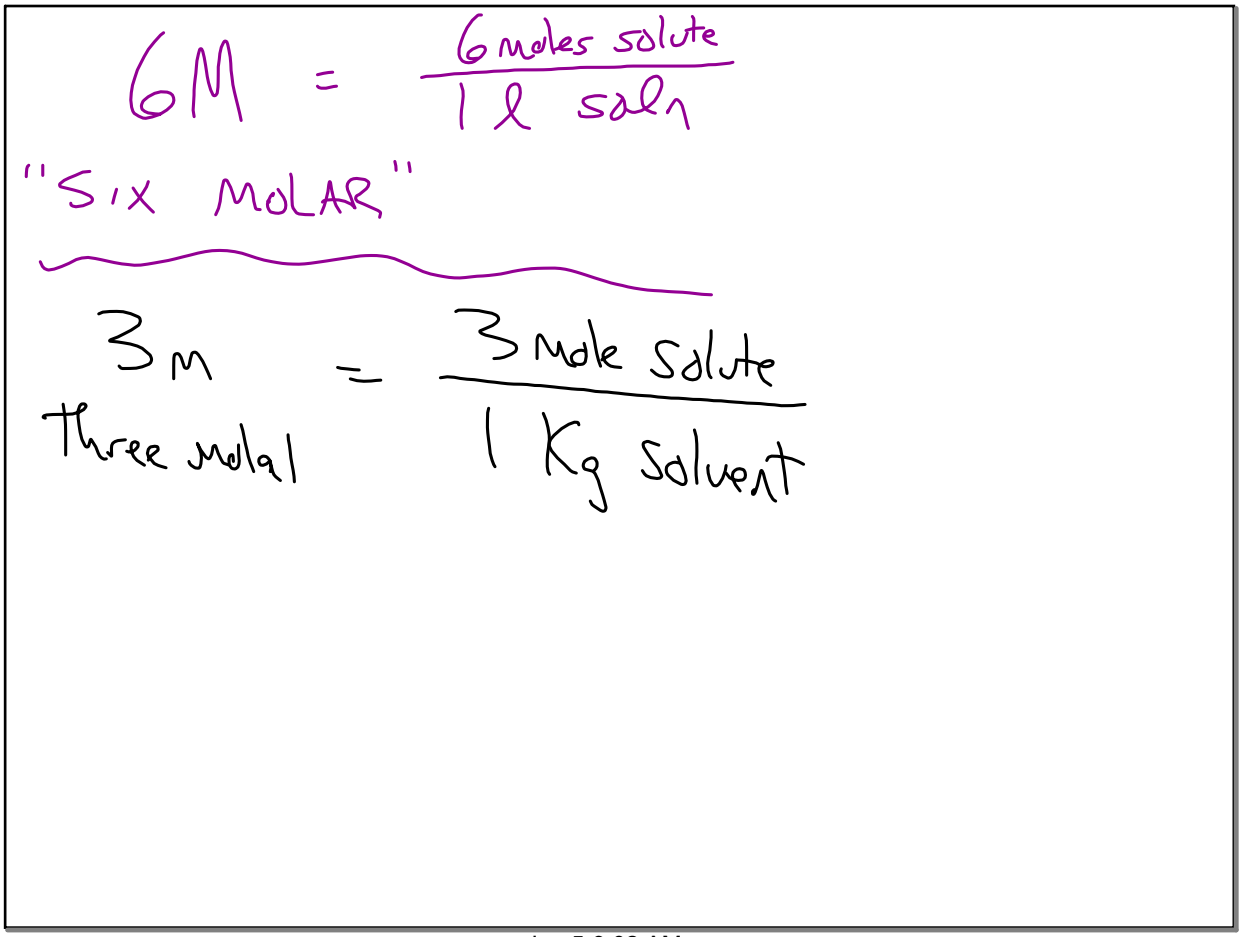
solution = solute + solvent

MORE

Jan 5-8:56 AM



Jan 5-8:58 AM



Jan 5-9:02 AM

Solute → 36.5 g $C_{10}H_8$ (Napthalene)
in Solvent → 425 g Toluene C_7H_8. D = 1 g/mL

Find ① molality + ② Molarity
↳ Find moles Napthalene / Kg Toluene

① $\frac{36.5\ g\ Nap}{0.425\ Kg\ Tol} \left| \frac{1\ Mole\ Nap}{128\ g\ Nap} \right. = 0.670_m$ $\frac{mole}{Kg}$ N T

Jan 5-9:03 AM

② Molarity = moles Nap / L solution
↳ N + T

$\frac{36.5\ g\ Nap}{461.5\ g\ soln} \left| \frac{1\ Mole\ Nap}{128\ g\ Nap} \right| \frac{1\ g}{1\ mL} \left| \frac{1000\ mL}{1\ L} \right.$

0.617 M

Jan 5-9:13 AM

13/28, 32, 44 c+d, 50a

Jan 5-9:16 AM