

(32) $O_2(g)$ $NaCl(aq)$
 (50) 1.5L 0.110M $(NH_4)_2SO_4$
 SOLUTION
 Find g $(NH_4)_2SO_4$

0.110 moles $(NH_4)_2SO_4$
 1 L solution

M

0.110 moles $(NH_4)_2SO_4$	1.5L	132 g $(NH_4)_2SO_4$
1L		moles $(NH_4)_2SO_4$

Jan 6-7:37 AM

Henry's Law

Gas escapes
↓
less soluble

Solubility of Gas = K P

↑
constant

Pressure on Gas

Soda
Open can
release P

Jan 6-8:03 AM

$$P_{\text{gas \#1}} = X_{\text{gas \#1}} P_{\text{TOTAL}}$$

$$\text{Mole fraction} = \frac{\text{Moles Part}}{\text{Total Moles}}$$

Jan 6-8:15 AM

Colligative Properties

Add solute to solvent :

ΔMP ~~FP~~ \downarrow and ΔBP \uparrow

Of the solution

Jan 6-8:17 AM

Name FP 0°C
Salt Cold Fusion (US) -4°C outside Temp
-10°C new FP

Jan 6-8:21 AM

- ① FP ↓
- ② BP ↑
- ③ Osmotic Pressure ↑ "Thicker" More concentrated

membrane
 $P = nRT$
 $\pi = iMRT$
 $P = nRT$
 $P = \frac{n}{V}RT$
 $P = MRT$

Jan 6-8:23 AM

$$13 / 42 + 46$$

Jan 6-8:28 AM