

(10.66) 10.25g N_2 , 1.83g H_2 , 7.95g NH_3
 0.366 mole 0.915 mole 0.468 mole

$P_T = 1.85 \text{ atm}$

$P_{N_2} = \underline{\hspace{2cm}}$
 $P_{H_2} = \underline{\hspace{2cm}}$
 $P_{NH_3} = \underline{\hspace{2cm}}$

$P = X_{\text{part}} P_T$
 (Hz) (Hz)

$n_T = 1.749 \text{ mole}$

$P_{N_2} = \frac{0.366}{1.749} (1.85) =$
 $P_{H_2} = \frac{0.915}{1.749} (1.85) =$
 $P_{NH_3} = \frac{0.468}{1.749} (1.85) =$

Dec 9-3:34 PM

(10.70) 3g SO_2 = 0.047 mole SO_2 2.35g N_2 = 0.084 mole N_2

~~V 5L~~ ~~20°C~~ ~~2.5L~~ ~~20°C~~
 T 21°C 10L 26°C V ?
 T ? T ?

$P_{SO_2} = \frac{nRT}{V}$
 $P_{SO_2} = \frac{0.047(0.08206)(299)}{10}$
 $P_{SO_2} = 0.115 \text{ atm}$

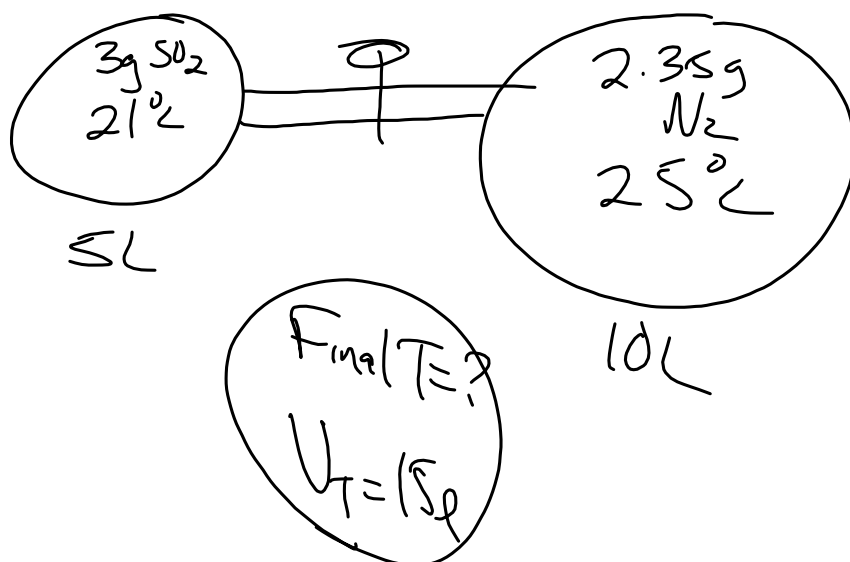
$P_{N_2} = \frac{nRT}{V}$
 $P_{N_2} = \frac{0.084(0.08206)(299)}{10}$
 $P_{N_2} = 0.206 \text{ atm}$

Dec 10-7:54 AM

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

SAME GAS

Dec 10-8:01 AM



Dec 10-8:02 AM

Kinetic Molecular Theory

Theory of molecules moving.

- "FAT GUYS RUN SLOW"
- Greater IMF (Van der Waals)

Elastic Collisions

Brownian Motion

Dec 10-8:05 AM

effusion

diffusion

Dec 10-8:21 AM