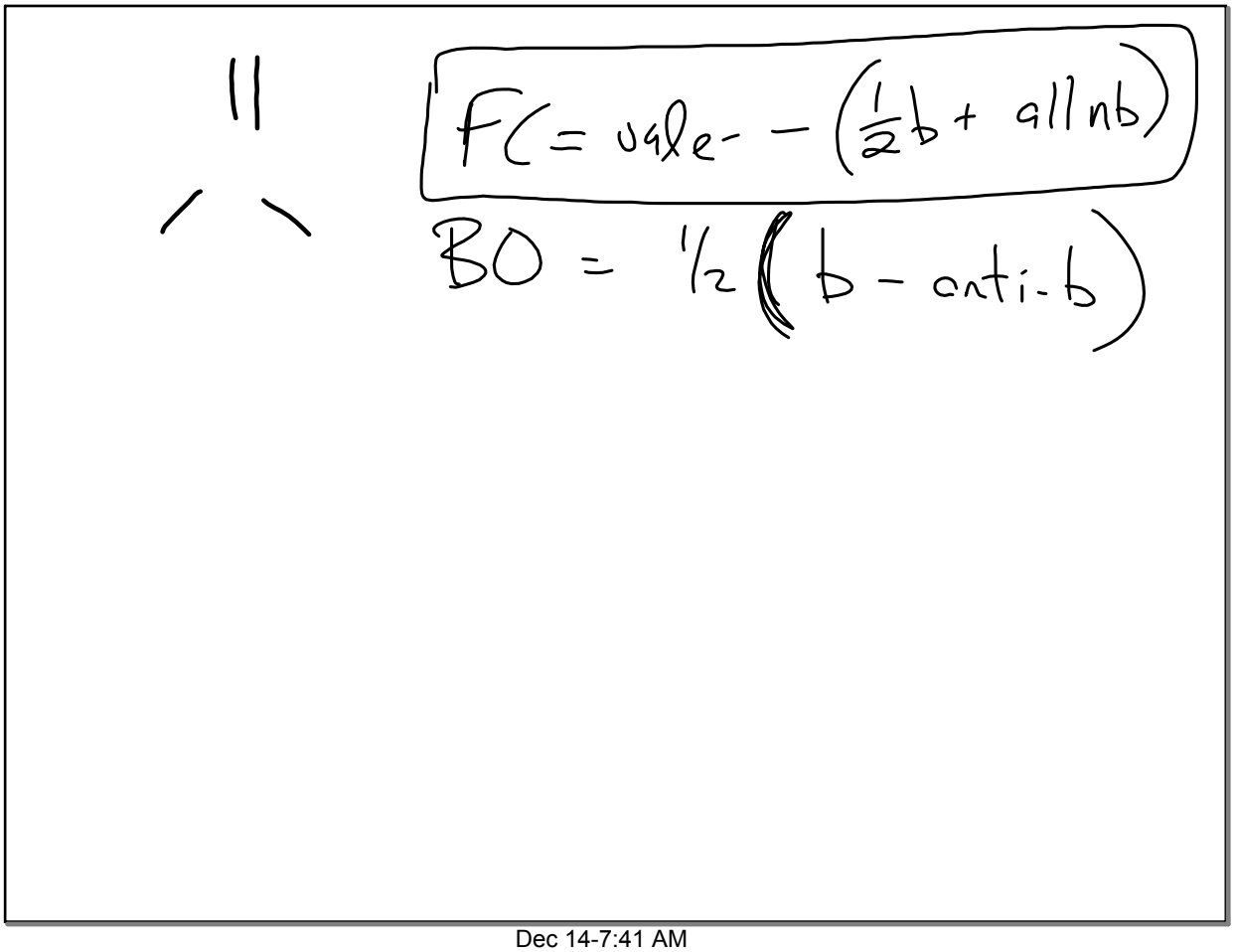
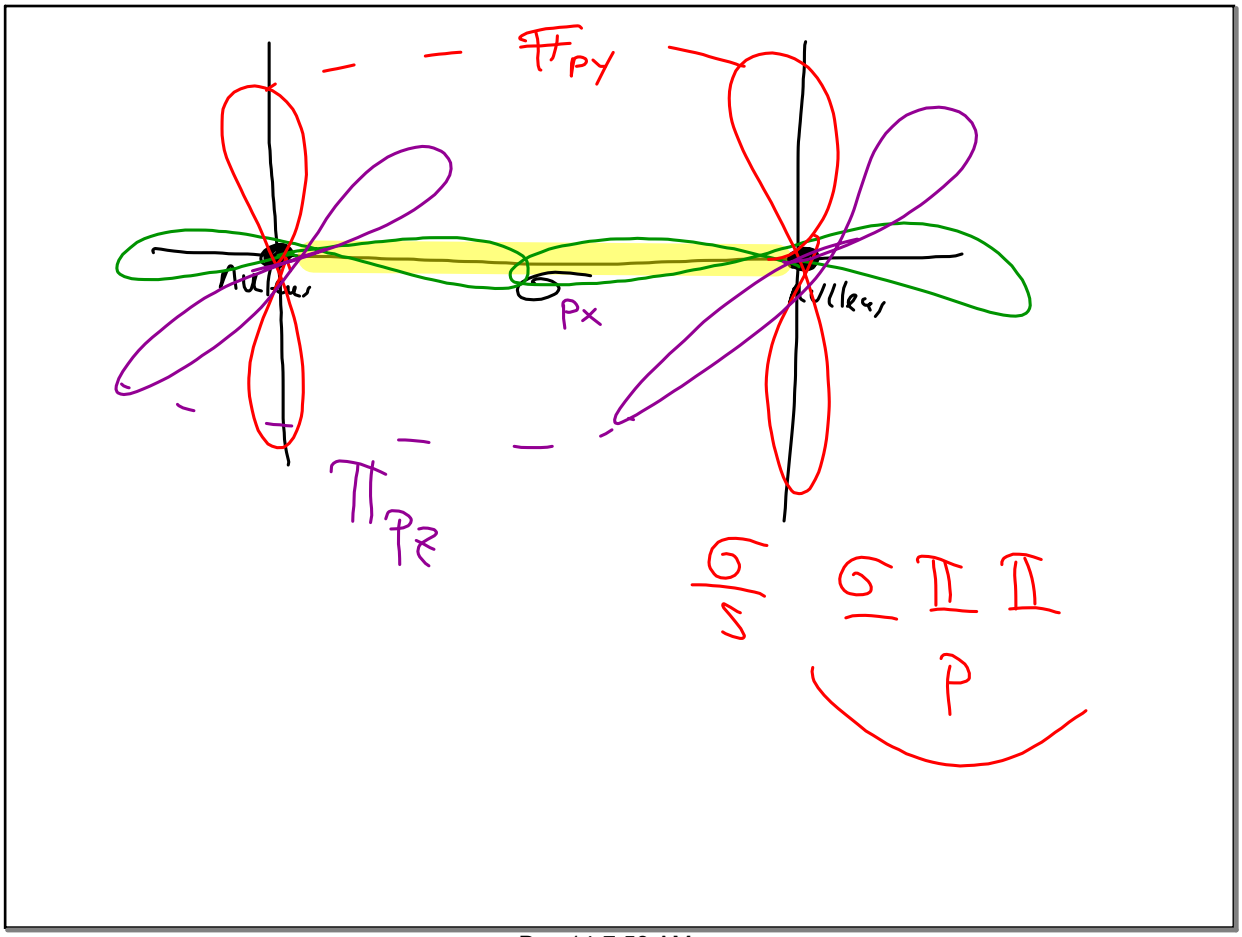


Dec 14-7:41 AM



Dec 14-7:58 AM

C_2H_2

$2(4) + 2(1) = 10 - 6 = 4$

H — C ≡ C — H

σ σ σ

2π

Dec 14-8:02 AM

(24)

Tank 1

V = 40 l

P = 1 atm

T = 20°C

Tank 2

V = 60 L

P = 2 atm

T = 20°C

$PV = nRT$

$n = \frac{PV}{RT}$

$V_T = 100$ l

$P_T V_T = n_T RT$

$P_T(100) = (0.8)(273)$

$n_2 = \frac{PV}{RT}$

Dec 14-8:12 AM

(29) $2NaN_3 \rightarrow 2Na + 3N_2$

— g

V= 40 ℓ

T= 25° = 298K

P= 763 mmHg · 1.00 atm/760 mmHg

$n = \frac{PV}{RT} = 1.64 \frac{mole}{}$

$\frac{1\ mole\ N_2}{3\ mole\ N_2} \quad \frac{2\ mole\ NaN_3}{3\ mole\ N_2} \quad \frac{65\ g\ NaN_3}{1\ mole\ NaN_3}$

Dec 14-8:17 AM

(30) $PV = nRT$

$\frac{PV}{1} = \frac{gRT}{MW}$

$g = \frac{PV(MW)}{RT}$

Dec 14-8:20 AM

(32) $d = \frac{5.75\,g}{\ell}$ at STP

273 K, 1 atm

0°C

$P_{gas} = X_{gas} P_T$

$= (0.018)(760)$

$= 13.68$ atm

Dec 14-8:21 AM