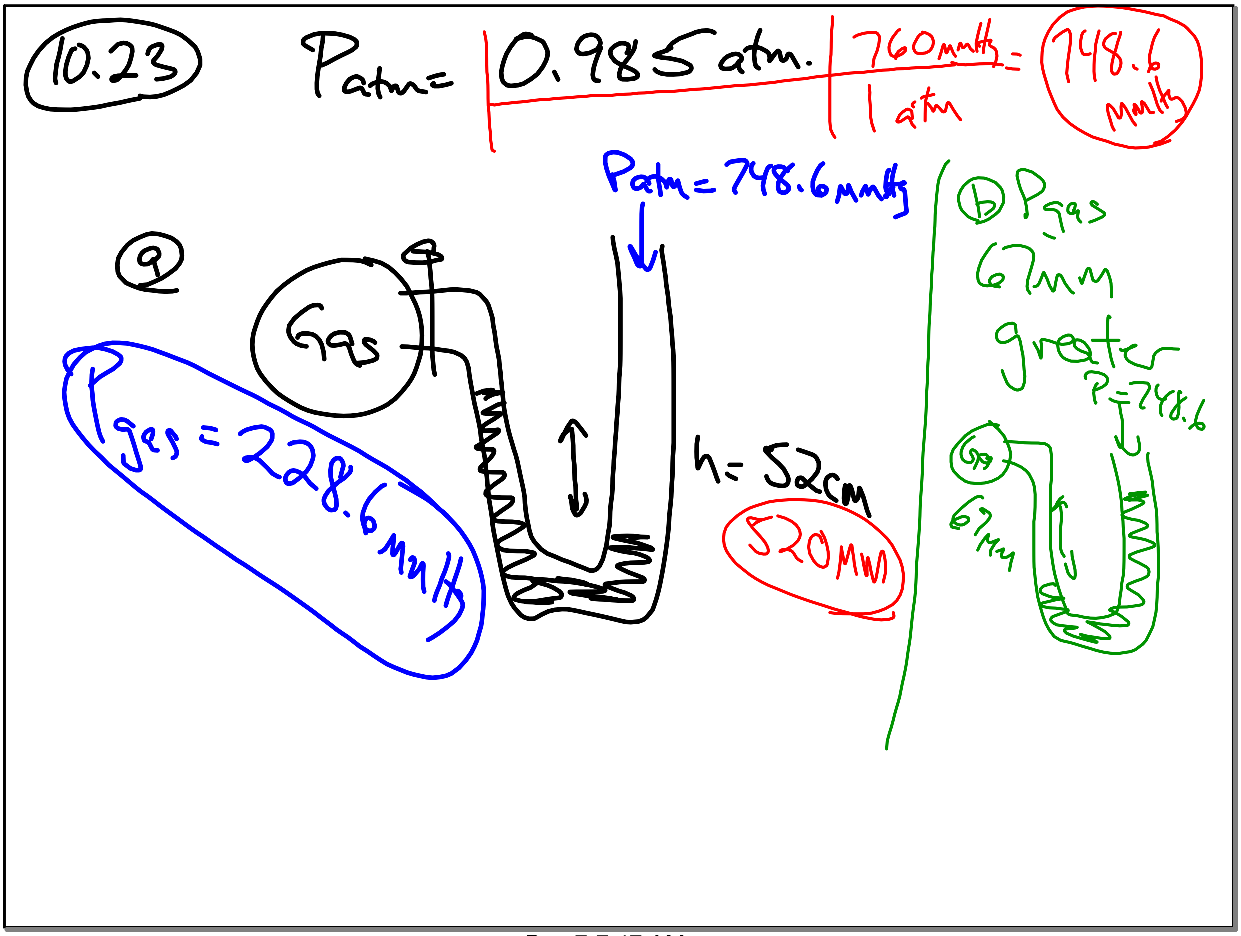
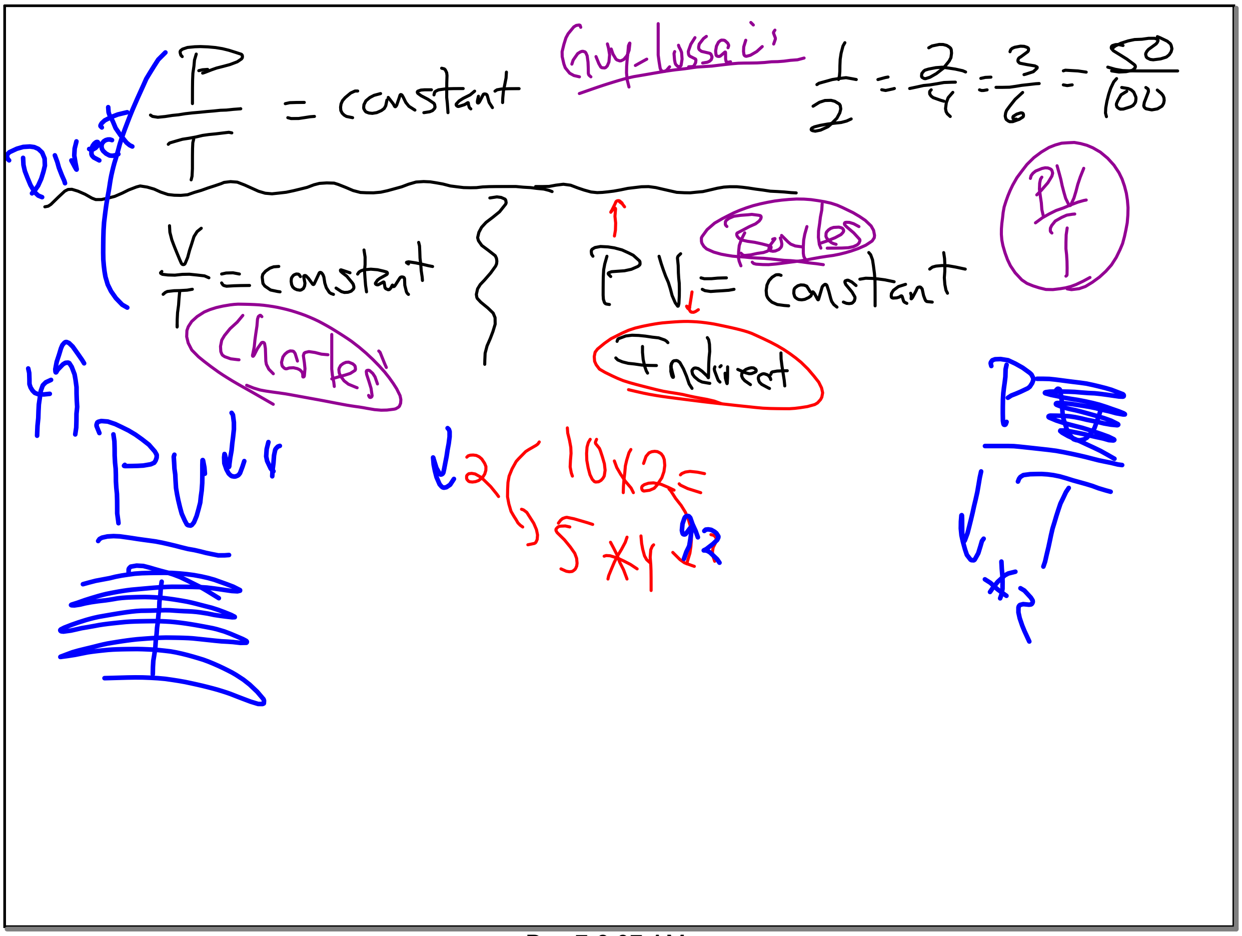


Dec 7-7:47 AM



Dec 7-8:07 AM

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

TEMP MUST BE IN KELVIN

Change in a system

Dec 7-8:21 AM

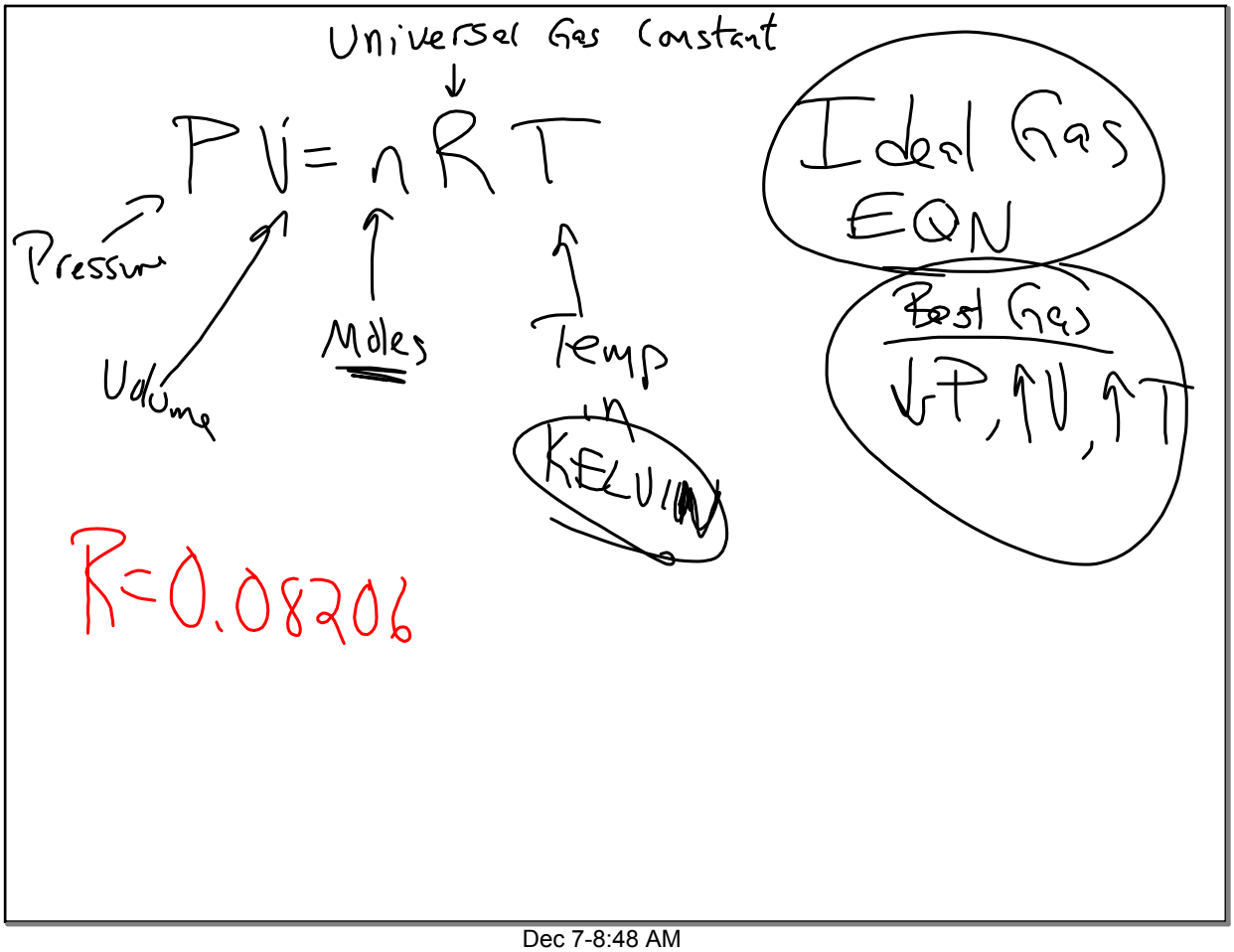
$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$\frac{252}{0.45}\frac{(1)(6)}{295} = \frac{(0.45)(V_2)}{252}$$

22°

$V_2 = 11.39\ l$

Dec 7-8:42 AM



Dec 7-8:48 AM

$PV = nRT$

$$R = \frac{PV}{nT} = \frac{\ell \cdot atm}{mole \cdot K} = 0.08206$$

Find P when. ⇒ 22g of CO_2 in 10 ℓ container at 300K

$$P = \frac{nRT}{V} = \frac{(0.5)(0.08206)(300)}{10}$$

$P = 1.231 atm$

$$\frac{\text{Moles}}{1} = \frac{g}{MW}$$

Dec 7-8:53 AM

Find molar mass

$PV = nRT$

$\frac{n}{1} = \frac{g}{MW}$

$\frac{PV}{1} = \frac{gRT}{MW}$

$\frac{MW}{1} = \frac{gRT}{PV}$

Dec 7-9:01 AM

Find Density of a gas

$D = \frac{\text{mass}}{\text{volume}} = \frac{g}{V}$

$D =$

$PV = nRT$

$\frac{PV}{1} = \frac{gRT}{MW}$

$\frac{P(MW)}{RT} = \frac{g}{V} = \text{density}$

Dec 7-9:03 AM

$2NaN_3 \rightarrow 2Na + 3N_{2(g)}$

— g

36 L

1.15 g/m

26 °C

$PV = nRT$

$n = \frac{PV}{RT}$

Dec 7-9:13 AM

10/34 a+b, 38, 40

Dec 7-9:17 AM