

(10.24a)

$0.966 \text{ atm} \times \frac{760 \text{ mm}}{1 \text{ atm}} = 734.16 \text{ mm}$
 734.16
 $- 15.4$

GAS
 15.4 mm
 718.76 mmHg

Dec 5-9:39 AM

(10.24b)

$0.99 \text{ atm} \times \frac{760}{1 \text{ atm}} = 752.4 \text{ mmHg}$
 752.4
 $+ 8.7$

GAS
 8.7 mm
 8.7 mm

Dec 5-9:54 AM

10.50 g 707 torr $\times \frac{1 \text{ atm}}{760 \text{ torr}}$ Find density $\frac{\text{g}}{\text{L}}$
 21°C
 SF_6

$PV = nRT$
 $\frac{P \cdot V}{RT} = \frac{g}{\text{MW}}$

$\frac{g}{L} = \frac{P(\text{MW})}{RT}$

$\frac{P(\text{MW})}{RT} = \frac{0.93(146)}{0.08206(294)}$
 5.628 g/L

Dec 5-9:57 AM

Parts is Parts is Pieces is Parts

$P_{\text{TOTAL}} = P_{\text{CO}_2} + P_{\text{H}_2\text{O}} + P_{\text{N}_2} + P_{\text{O}_2}$
 Dalton

Dec 5-10:01 AM

$P_{O_2} = ?$

MOLE fraction

$P_{O_2} = X_{O_2} P$

Part / Whole = Molar O_2 / Total Moles in container

Dec 5-10:08 AM

10/43, 62a

Dec 5-10:11 AM