

2016 (12) $\text{KOH} + \text{HCN} \rightarrow \text{KCN} + \text{H}_2\text{O}$

20ml, 0.3M 100ml, 0.15M
 0.006 moles 0.015 moles

Moles = M * Q

* Mix A + B \Rightarrow Next. MOLES!

Q Next moles

$\text{KOH} + \text{HCN} \rightarrow \text{CN}^- + \text{H}_2\text{O}$	Q	Q	Q
I 0.006 0.015	-0.006	-0.006	+0.006
E Q	0.009	0.009	0.006
	0.12L	0.12L	0.12L
	0.75M HCN	0.75M KCN	0.05M KCN

Q Realc. new M

HH

$$\text{pH} = \text{pK}_a + \log \frac{[\text{base}]}{[\text{acid}]} = \text{pK}_a + \log \frac{[\text{CN}^-]}{[\text{HCN}]}$$

May 2-7:42 AM

(13) $A_0 = 0.26 \text{ M}$

(2°) $A_{\infty} = 0.13 \text{ M}$

$k = 0.13 \frac{1}{\text{M} \cdot \text{Sec.}}$

$$\frac{1}{A_t} = kt + \frac{1}{A_0}$$

May 2-8:03 AM

(17)

$$\text{pH} + \text{pOH} = 14$$

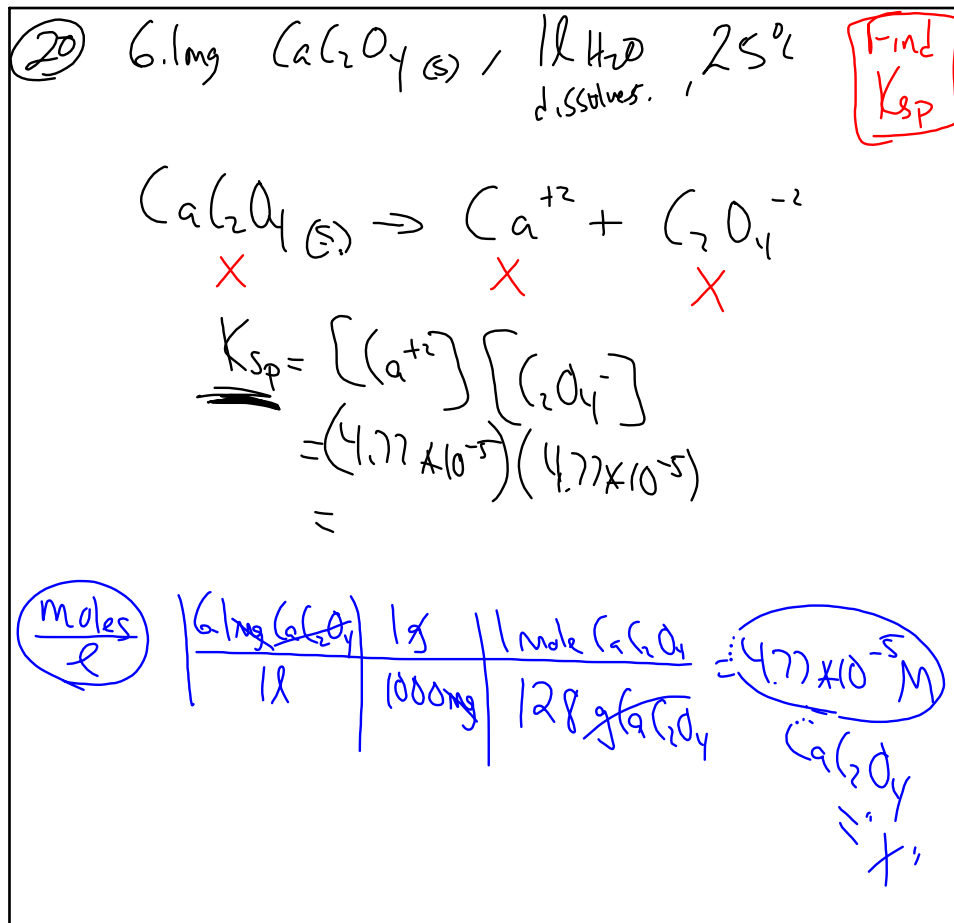
$\text{pH} = 4.282$
 $\text{pOH} = 9.718$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$9.718 = -\log[\text{OH}^-]$$

$$-9.718 = \log[\text{OH}^-]$$

May 2-8:07 AM



May 2-8:10 AM

Tarea para esta noche
2015 Final #1-20

May 2-8:26 AM