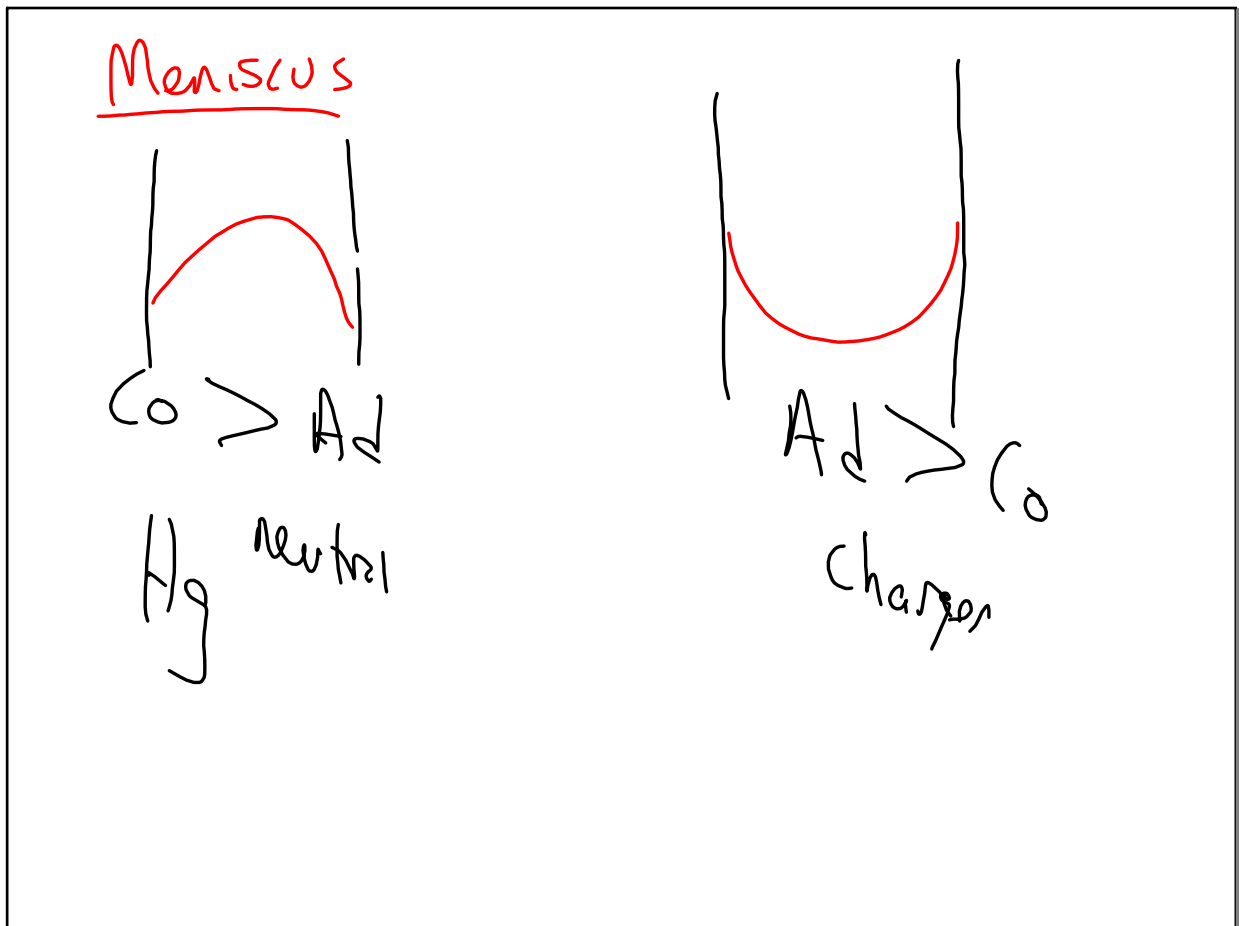


Jan 2-7:37 AM



Jan 2-8:05 AM

②②

Temp °C

Ice (s) at 0°C

16.7 kJ H₂O

53

tm

① $Q = mH_f$
 $\frac{6.01 \text{ kJ/mol}}{18 \text{ g/mol}} \times 50 \text{ g} = 16.7 \text{ kJ}$

② $Q = mc\Delta T$

75.2 J	1 mol	50 g	22 K
mol	18 g		

C M ΔT 4.6 kJ

21.3 kJ

Jan 2-8:09 AM

②④

T

10 kJ 15.5g ice

-5°C → 69.5 Temp

10.00 0.32

9.68 kJ/mol

69.5

4.5 kJ

time

① -5°C to 0°C

75.2 J	5	mol	15.5g	= 0.32 kJ
mol	18 g			

②

6.01 kJ	1 mol	15.5g	= 5.18 kJ
mol	18 g		

③

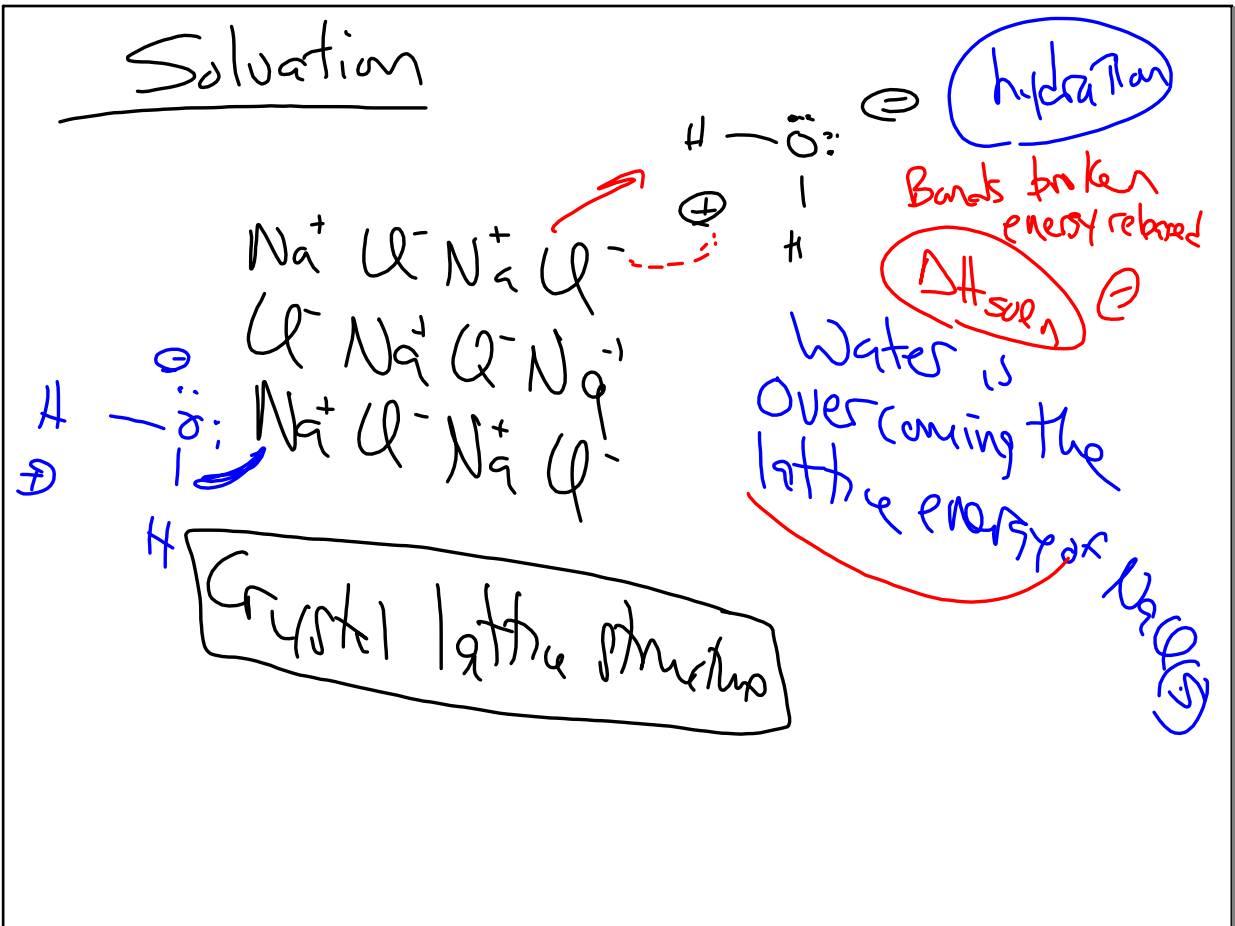
mol	4.5 kJ	18 g	15.5g	=
75.2 J				

69.5°C

Jan 2-8:19 AM

Chap 13
Solutions → homogeneous mixture.
Solvent → does the dissolving
Solute → gets dissolved.
 → If same phase (solute + solvent)
 Solute → lesser quantity.

Jan 2-8:46 AM



Jan 2-8:57 AM

Saturated - "Full"

All solvent seats are occupied
by a solute (MOVIE THEATER)

If more solute is added, that
same amount will precipitate out.

FACTORS Affecting ① Temp ② Pressure (GASES)

Jan 2-9:11 AM

UNSATURATED

13 / 24 + 28

Jan 2-9:15 AM