

⑧ At mass 116  
 p+n 116  
 # p → 92  
 At #

⑨ Ga<sup>+3</sup> X<sup>-3</sup>

Jan 29-7:41 AM

⑬ H<sub>2</sub>SO<sub>4</sub> moles A = moles B NaOH  
 $n \times M \times Q = n \times M \times Q$   
 (2)(0.0875)(Q) = (1)(0.115)(ml)  
 76.1ml

Jan 29-8:12 AM

(21)  $M = ?$  4.57 L of 0.847 M  $\text{Na}_3\text{PO}_4$

$\text{Na}^+$

$$\text{Na}_3\text{PO}_4 \rightarrow 3\text{Na}^+ + \text{PO}_4^{3-}$$

0.847 M - M

: \*3

3

Jan 29-8:16 AM

(23)  $\text{N}_2\text{O}_3 \rightarrow \text{N}_2 + \frac{3}{2}\text{O}_2$  -83.7

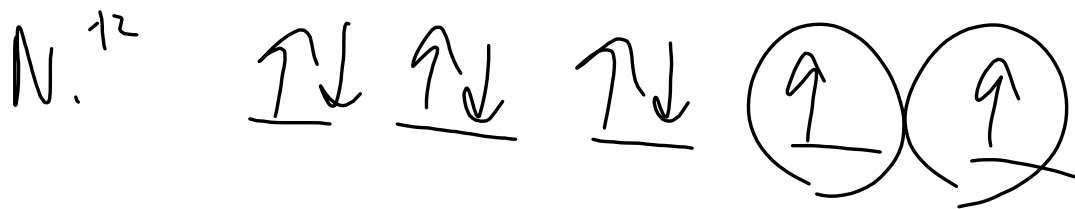
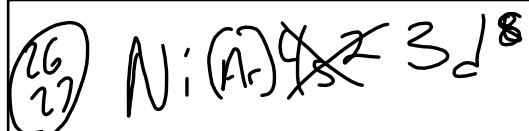
$\frac{1}{2}\text{N}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{NO}$   $\frac{1}{2}(180.4)$

$\frac{1}{2}\text{N}_2 + \text{O}_2 \rightarrow \text{NO}_2$  33.2

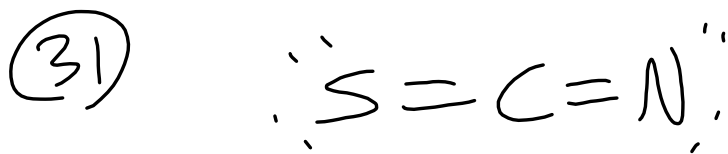
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+ 39.7 kJ

Jan 29-8:19 AM



Jan 29-8:23 AM



$$FC = \text{val}e^- - \left( \frac{1}{2}b + \text{all } nb \right)$$



Jan 29-8:35 AM

(23)

$$\frac{PV}{T} = \frac{PV}{T}$$

$$\frac{(100)(900)}{1} = \frac{P(300)}{2}$$

Jan 29-8:38 AM

(E1)

$$\text{CaCO}_3 + 2\text{HF} \rightarrow \text{CaF}_2 + \text{H}_2\text{O} + \text{CO}_2$$

50g (100ml / 1M) 0.1M — mols

MxR  
0.1mole → 0.05mole

50g CaCO <sub>3</sub>	1mole CaCO <sub>3</sub>	1mole CO <sub>2</sub>	=	0.5mole
100g CaCO <sub>3</sub>	1mole CaCO <sub>3</sub>	1mole CO <sub>2</sub>		

Jan 29-8:39 AM

$(Fe)$  Heat lost = Heat gained  $(Water)$   
 $mC \Delta T = mC \Delta T$   
 $(95)(C)(75-23) = (50)(4.184)(23-18)$   
 $C = 0.211 \text{ J/g}^\circ\text{C}$

Jan 29-8:43 AM

FALL CHE 106 / 107  
 Lec (3) Lab (1)  
 Spring 116 / 117

Jan 29-8:49 AM

# Chap 14 - Chemical Kinetics

how a chemical rxn proceeds.

**RATE** of a rxn.

↳ how fast

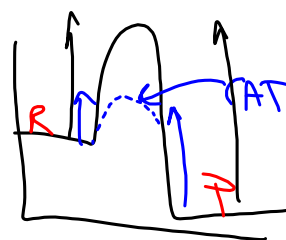
$$\frac{dx}{dt} = \frac{\text{change in } X}{\text{change in time}}$$

Jan 29-8:51 AM

## Rate determining step

Factors → head on collision - FAST

- ① Temp → Avg KE
- ② Concentration  $M = \frac{\text{Moles}}{V}$  [ ]  
↳ high [ ]
- ③ ↓ Volume
- ④ Catalyst ↓  $E_a$
- ⑤ ↑ Surface Area



Jan 29-8:58 AM

	$\xrightarrow{\quad}$	
$\uparrow$ 7 AM time $\&$	Reactants 100%	Products $\&$ 0%
	$\downarrow \Delta 90$	$\downarrow \Delta 90$
3 pm End time	10%	90%
	Rate of disappearing Reactants	Rate of appearance Products

Jan 29-9:09 AM

Efficiency Curve

$$\frac{d[\ ]}{dt} = \frac{\Delta[\ ]}{\Delta t}$$

Jan 29-9:12 AM



$$-\frac{1}{2} \frac{\Delta[A]}{\Delta t} = -\frac{1}{3} \frac{\Delta[B]}{\Delta t} = +\frac{1}{1} \frac{\Delta[C]}{\Delta t}$$

$\frac{1}{\text{coeff}}$  (RATE)

$$14 / 18 + 20$$

Jan 29-9:14 AM