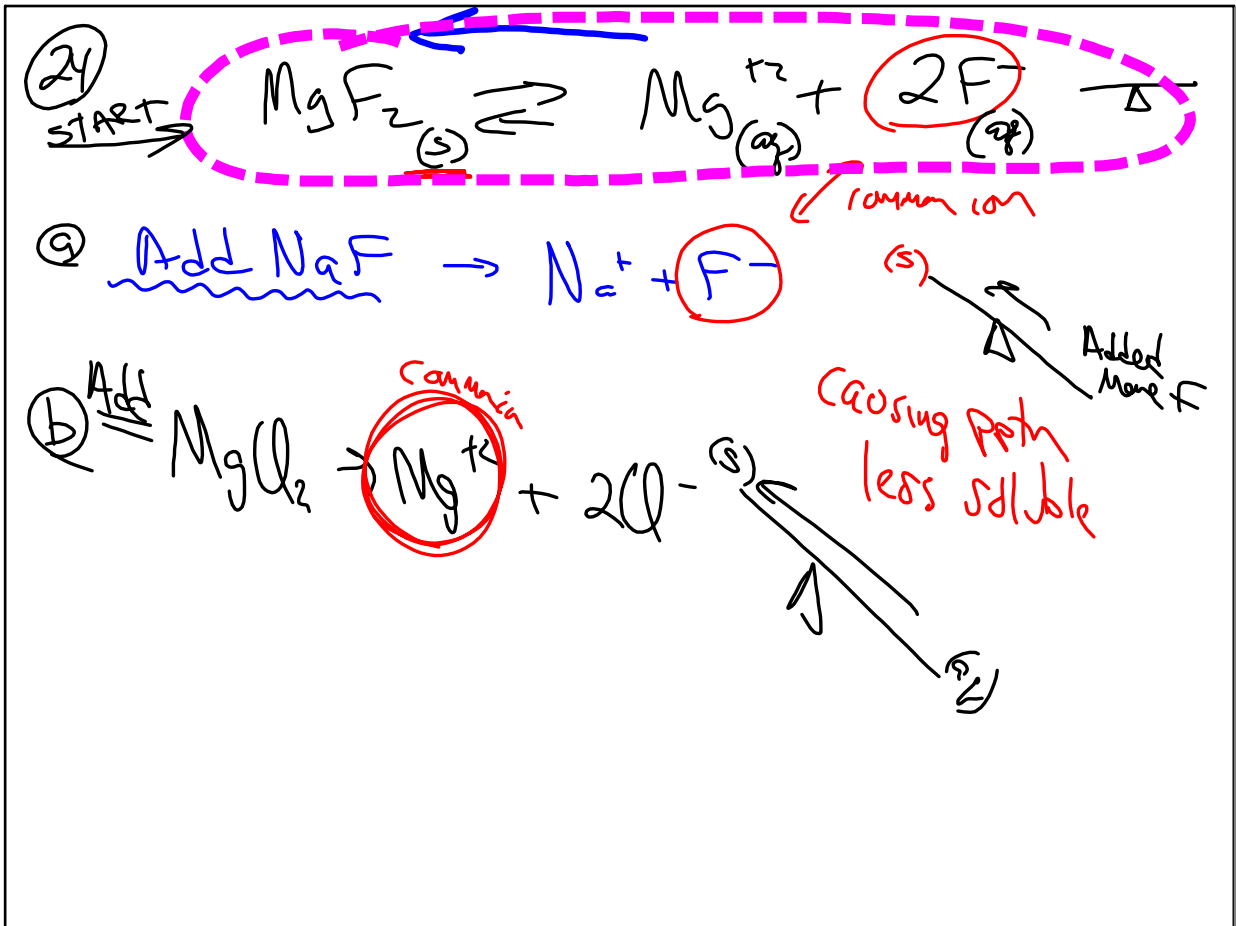
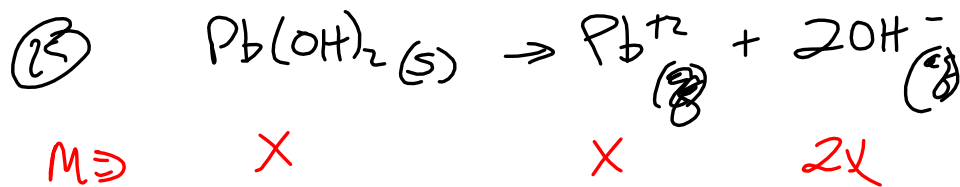


Mar 16-8:23 AM



Mar 16-8:46 AM



$$K_{sp} = [\text{Pb}^{+2}] [\text{OH}^-]^2$$

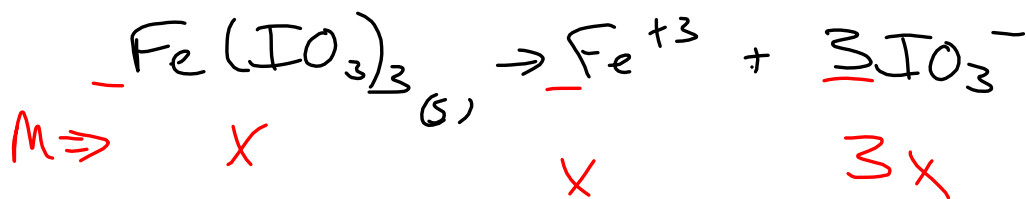
$$2.4 \times 10^{-16} = (x) (2x)^2$$

$$x = 3.91 \times 10^{-6}$$

Mar 16-8:53 AM

(26) $K_{sp} = 10^{-49}$

$Q < K$



$$Q_{sp} = [\text{Fe}^{+3}] [\text{IO}_3^-]^3$$

$$= (10^{-4}) (10^{-5})^3$$

$Q_{sp} = 10^{-19}$

Mar 16-8:57 AM

(27)

moles A = moles B

$$n_A \times M_A \times l_A = n_B \times M_B \times l_B$$

Mar 16-9:01 AM

(E2) Find pH

0.005 mol
50 ml 0.1M NaOH + $K_a = 6.8 \times 10^{-4}$
30 ml 0.25M HF
0.0075 moles.

① NaOH with moles

	NaOH ⁻	+	HF	→	NaF ⁻	+	H ₂ O
I	0.005		0.0075		0		
Δ	-0.005		-0.005		+0.005		
F	0		0.0025		0.005		

0.08 l
0.03125M HF

0.08 l
0.0625M F⁻

② Recalc. new M
80 ml 0.080

③
$$pH = pK_a + \log \frac{b}{a}$$

$$= -\log(6.8 \times 10^{-4}) + \log \frac{0.0625}{0.03125}$$

$$= 3.47$$

Mar 16-9:04 AM