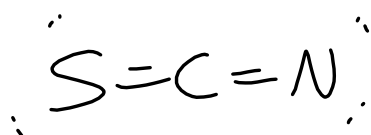


Exam 3

③



$$\text{Val e}^- - \left(\frac{1}{2} \text{b} + \text{all nb} \right)$$

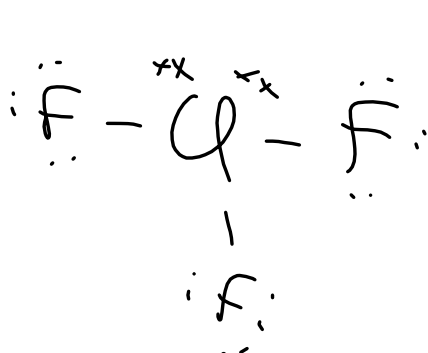
④ NO⁻

$$5 + 6 + 1 = 12$$

Jan 15-7:26 AM

⑥ CF₃

$$7 + 3(7) = 28 - 6 = 22 - 18 = 4 - 2 = 2$$

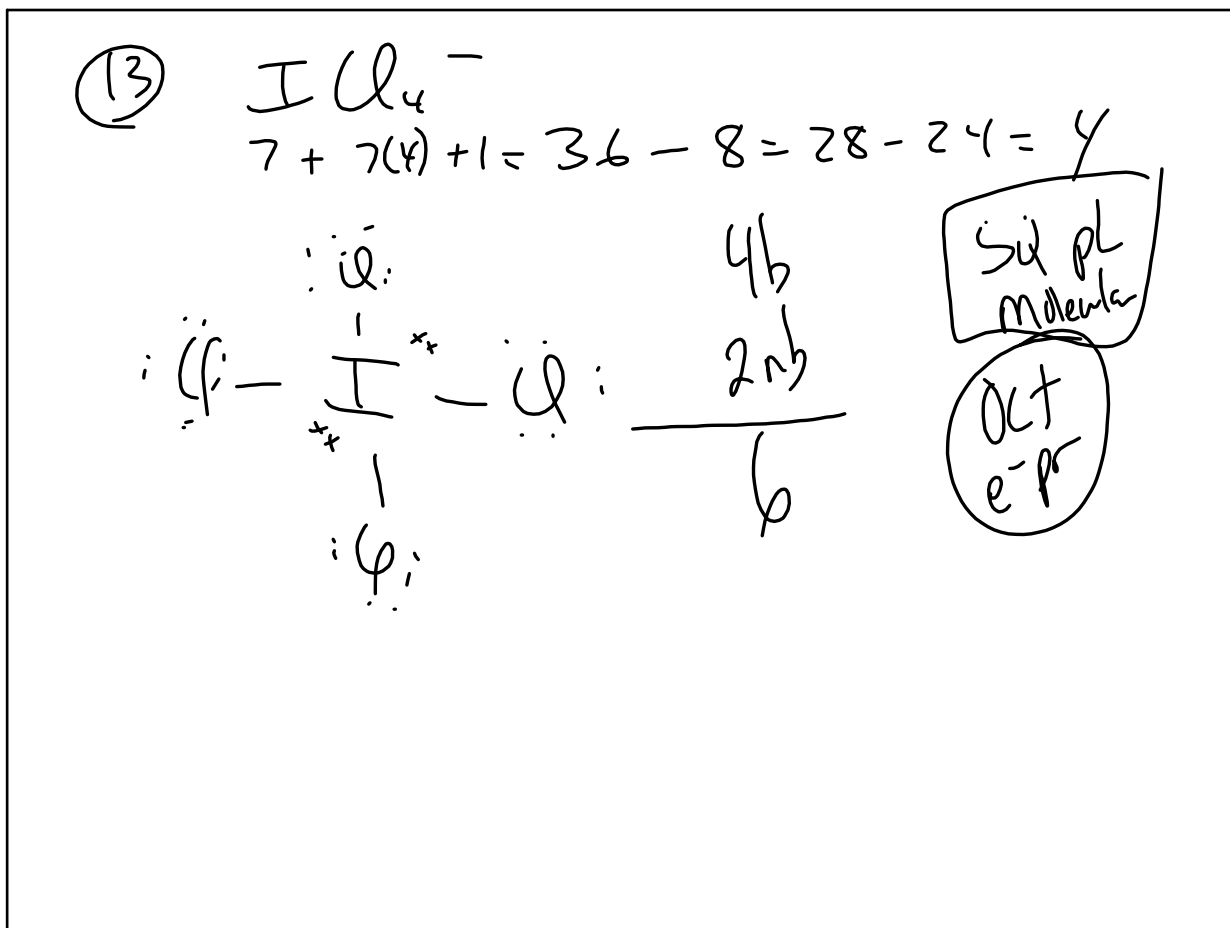
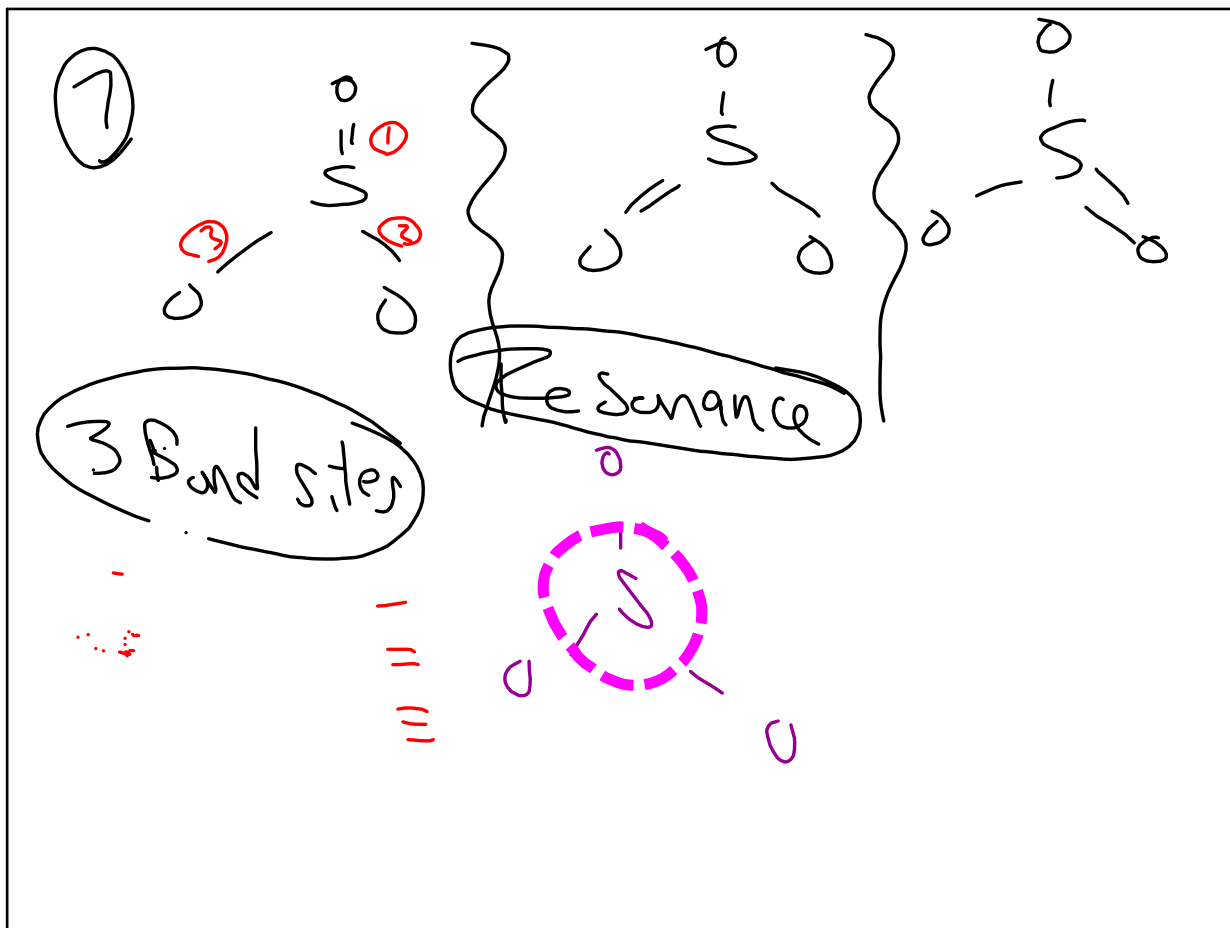


$$\text{sp}^3 \downarrow$$

⑤

$$\frac{-2}{\cancel{2}}$$

Jan 15-7:34 AM



$$(19) \quad PV = nRT$$

$$d = \frac{3.48 \text{ g}}{l}$$

$$\frac{PV}{l} = \frac{g RT}{MW}$$

$$MW = ? \quad \frac{g}{mol}$$

$$MW = \frac{g RT}{PV} = \frac{3.48 (0.08206) 273}{(1) (22.4)}$$

$$= \frac{3.48 (0.08206) (273)}{1}$$

Jan 15-7:56 AM

(23)

$$\frac{PV}{T} = \frac{PV}{T}$$

$$\frac{(100)(900)}{1} = \frac{P (300)}{2}$$

Jan 15-8:10 AM

(24)
$$P_{Ne} = X_{Ne} P_T$$

$$\frac{4}{10} (1000)$$

Jan 15-8:15 AM

(EC) $\rightarrow V = 23.88 \text{ L}$

$$1 \text{ NH}_4\text{NO}_2 \text{ (s)} \rightarrow \text{N}_2 \text{ (g)} + 2\text{H}_2\text{O (g)}$$

35g = 0.547 mole
 525°C → 798K
 1.5 atm
3 mole gas!

35g NH ₄ NO ₂	1 mole NH ₄ NO ₂	= 0.547 mole
	64g NH ₄ NO ₂	

$PV = nRT$

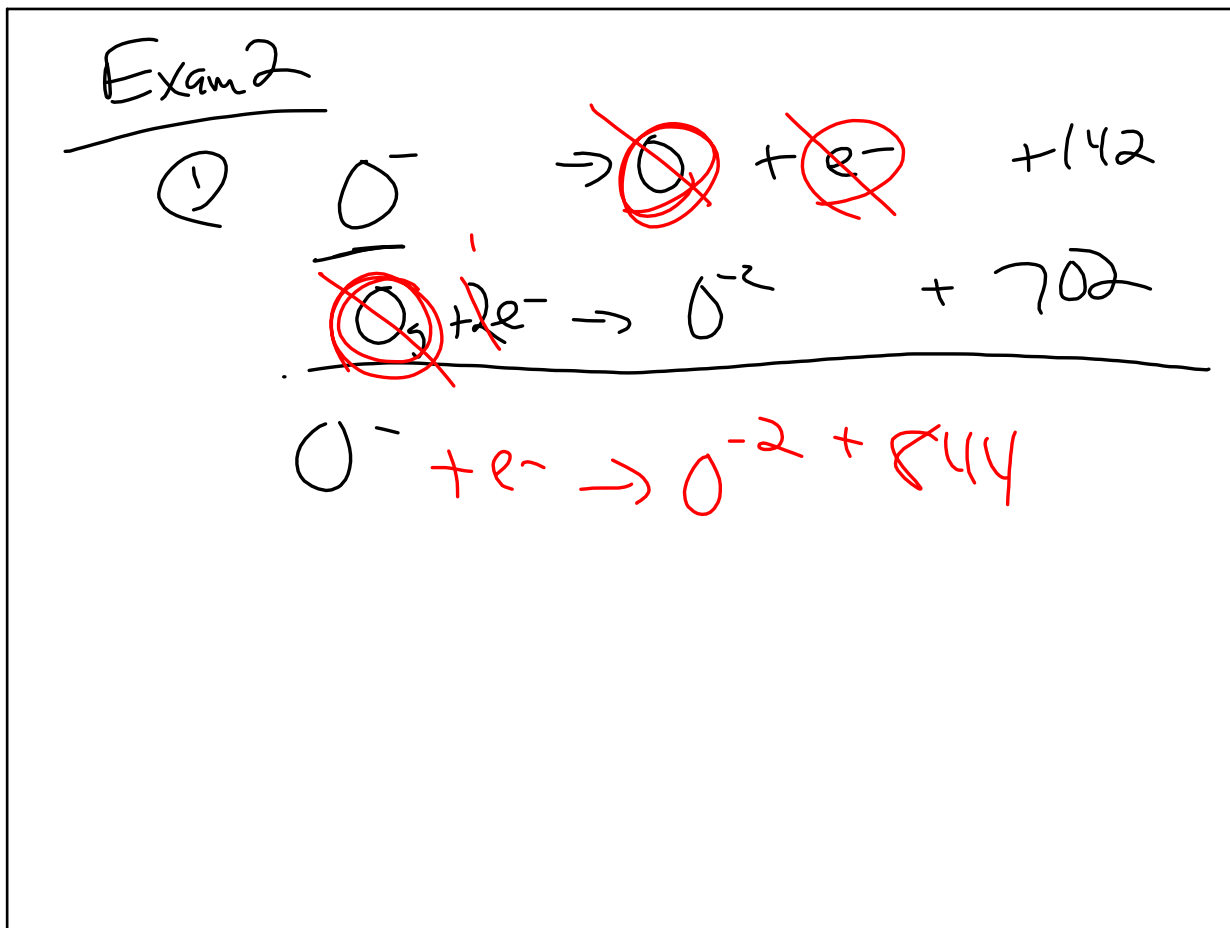
$$V = \frac{nRT}{P} = \frac{(0.547)(0.08206)(798)}{1.5}$$

$V = 23.88 \text{ L}$ (1 mole)

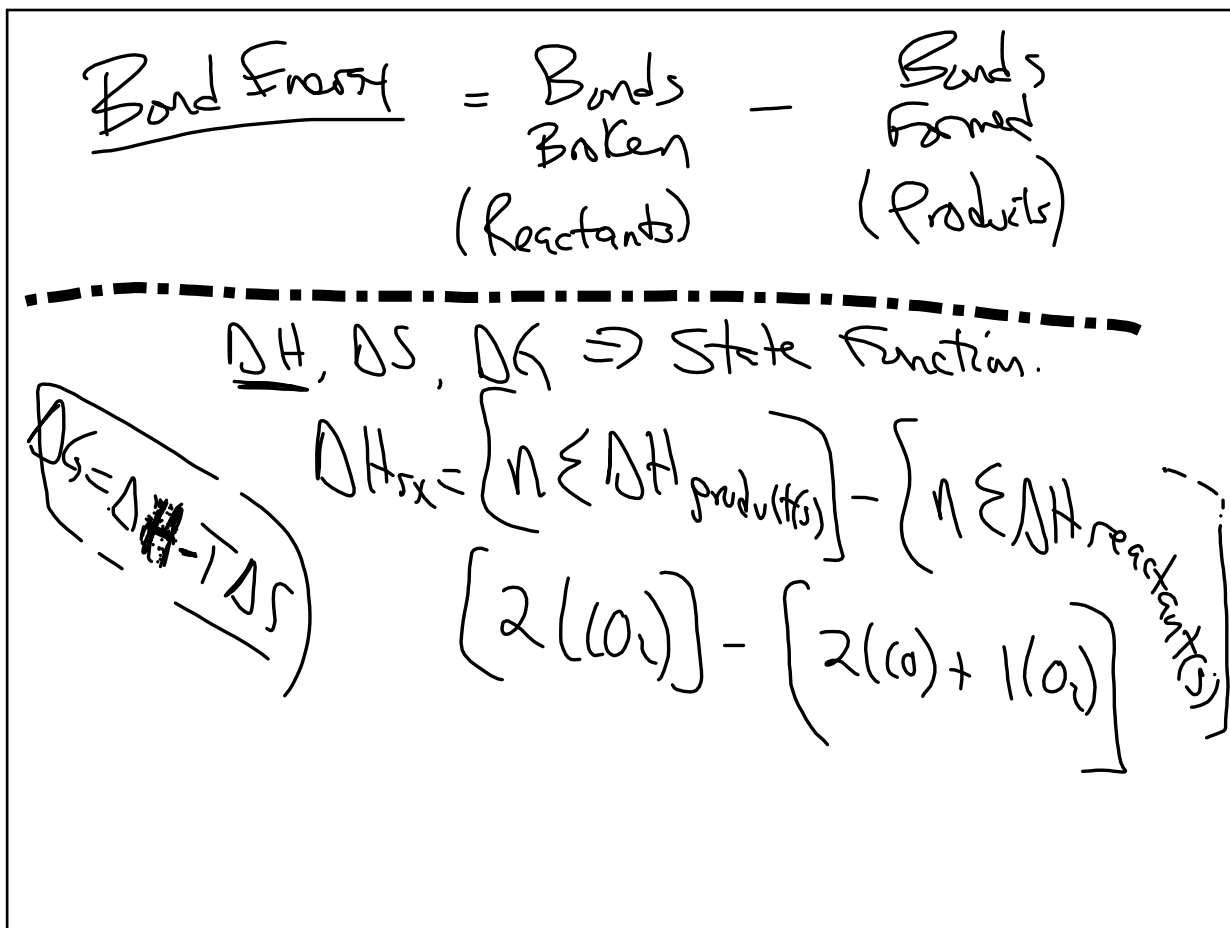
$\times 3$ (3 mole)

71.64 L

Jan 15-8:19 AM



Jan 15-8:25 AM



Jan 15-8:31 AM

$\frac{\uparrow}{P} \quad \text{---} \quad \text{---}$

Jan 15-8:36 AM

(B) $\lambda = 280 \text{ nm}$, $f = ?$
 $c = \frac{3 \times 10^8 \text{ m}}{\text{sec}}$
 $c = f \lambda$
 $3 \times 10^8 = f (280 \times 10^{-9} \text{ m})$
 $1.07 \times 10^{15} \text{ sec}^{-1}$

Jan 15-8:41 AM